

Preparing for the Organic Chemistry Exam?

1. Basic Concepts

Atomic Structure: Atoms are made of three sub-particles: Electrons, Protons, and Neutrons.

Electron: Negatively charged particle, mass 9.1×10^{-31} kg.

Proton: Positively charged particle, mass 1.67×10^{-27} kg.

Neutron: Neutral particle, mass 1.67×10^{-27} kg.

Atomic Number (Z): Number of protons in the nucleus.

Mass Number (A): Sum of protons and neutrons in the nucleus.

Isotopes: Atoms of the same element with different mass numbers.

2. Atomic Structure of Matter

Atomic Model: The atom consists of a central nucleus surrounded by electrons.

Electron Configuration: Electrons are arranged in shells or orbitals.

Bohr's Model: Electrons move in fixed orbits around the nucleus.

Spectral Lines: Discrete lines of light emitted or absorbed by atoms.

3. Periodic Table

Periodic Law: The properties of elements are a periodic function of their atomic number.

Groups: Vertical columns of elements with similar properties.

Periods: Horizontal rows of elements.

Metals: Elements on the left side of the periodic table.

Non-metals: Elements on the right side of the periodic table.

4. Chemical Bonding

Electron Dot Diagram: A diagram showing the valence electrons of atoms.

Valency: The combining capacity of an element.

Chemical Formula: A chemical symbol or formula representing a compound.

Molecular Formula: A chemical formula representing a molecule.

5. Chemical Reactions

Chemical Reaction: A process in which one or more substances are converted into one or more new substances.

Chemical Equation: A chemical equation representing a chemical reaction.

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6. Acids and Bases

Acid: A substance that releases H^+ ions in aqueous solution.

Base: A substance that releases OH^- ions in aqueous solution.

pH Scale: A scale used to measure the acidity or basicity of a solution.

Indicators: Substances that change color in acidic or basic solutions.

7. Metals and Non-metals

Metals: Elements that are lustrous, malleable, and good conductors of heat and electricity.

Non-metals: Elements that are not lustrous, brittle, and poor conductors of heat and electricity.

Alloys: A mixture of two or more metals.

Corrosion: The process of deterioration of a metal due to chemical reaction with the environment.

8. Carbon and its Compounds

Carbon: An element with atomic number 6.

Carbon Compounds: Compounds containing carbon.

Hydrocarbons: Compounds containing only carbon and hydrogen.

Functional Groups: Groups of atoms that determine the chemical properties of a compound.

9. Organic Chemistry

Organic Chemistry: The study of carbon compounds.

Organic Compounds: Compounds containing carbon.

Organic Reactions: Chemical reactions involving organic compounds.

Organic Synthesis: The process of synthesizing organic compounds.

10. Environmental Chemistry

Environmental Chemistry: The study of chemical processes in the environment.

Pollution: The introduction of harmful substances into the environment.

Acid Rain: Rainwater with a low pH.

Global Warming: The increase in the average temperature of the Earth's atmosphere.

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