

SUBJECT: CHEMISTRY

CLASS: FORM 2

TOPIC: 1, 2 & 3.

: STRUCTURE OF THE ATOM AND THE PERIODIC TABLE.

: CHEMICAL FAMILIES; PATTERNS AND PROPERTIES.

: STRUCTURE AND BONDING.

1. The electron arrangement of ions Q^{2-} and R^{3+} are as 2, 8, 8, and 2, 8 respectively.

(a) Write the electron arrangement of the elements Q and R

Q 2, 3, 6

R 2, 8, 3

(b) Write the formula of the compound that would be formed between Q and R

R_2Q_3

2. The 2.table below gives some properties of three elements X, Y and Z.

ELEMENT	Atomic No.	Meeting point($^{\circ}C$)	Boiling Point ($^{\circ}C$)
X	53	114	184
Y	35	-7	58.8
Z	17	-101	-34.7

(a) Which element is in liquid form at room temperature? Give reason.

Y-At room temperature (25 $^{\circ}C$) Y is a liquid since its M.P and B.P lie between -7 $^{\circ}C$ and 58.8 $^{\circ}C$ OR

(Room temperature is between M.P and B.P.)

(b) Explain why the boiling point of element X is higher than that of element Z.

The molecular mass of x is higher than that of Z; the Vander Waals forces are stronger ✓ in X molecules than in Z molecules hence X has a higher boiling point than Z.

3. The elements A and B have the following properties

Element	Mass No.	Atomic No.
A	37	17
B	37	18
C		

When the isotope A was bombarded with a neutron, an isotope C was formed .Fill in the table to show the properties of element C.

Mass No of C=38

Atomic No =17

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