

# The Mole Concept

## Counting by weighing

The size of molecule is so small that it is physically difficult if not impossible to directly count out molecules. this problem is solved using a common trick. Atoms and molecules are counted indirectly by weighing.

Here is a practical example. You need to estimate the number of nails in a box. You weigh an empty box, 213. g. The weight of the box plus nails is 1340. g. The weight of one nail is 0.450 g.

I hope you aren't going to tear open the package and count the nails. We agree that

$$\text{mass of nails} = 1340 \text{ g} - 213 \text{ g} = 1227. \text{ g}$$

$$\begin{aligned} \text{Number of nails} &= (1227. \text{ grams nails})(1 \text{ nail}/ 0.450 \text{ grams}) \\ &= 2726.6 \text{ nails} = \mathbf{2730 \text{ nails}} \end{aligned}$$

You can count the nails by weighing them.

## Avogadro's Number and the Mole

To calculate real chemical reactions in the laboratory, chemists use a special unit called a mole (abbreviated mol). One mole of a substance is the amount of the substance that is equal in molar mass of the molecular or formula mass of the substance in grams. Thus for ethylene,  $\text{C}_2\text{H}_4$ , its molecular weight is 28 and its molar mass is 28 g. In other words, 28 g represents 1 mol of ethylene. One mole contains  $6.022 \times 10^{23}$  molecules or formula units. The number  $6.022 \times 10^{23}$  is called Avogadro's number.

$$\begin{aligned} \text{For ethylene, } \text{C}_2\text{H}_4, \\ \text{Molecular mass } \text{C}_2\text{H}_4 &= 28. \\ \text{Molar mass } \text{C}_2\text{H}_4 &= 28.0 \text{ g} \\ 1 \text{ mol } \text{C}_2\text{H}_4 &\text{ contains } 6.022 \times 10^{23} \text{ molecules} \end{aligned}$$

The coefficients of a balanced chemical equation indicate the number of moles of each substance in the reaction. Thus, at the level of moles:



One mole of propane reacts with five moles of oxygen to form three moles of carbon dioxide and four moles of water.

Avogadro's number is an accident of nature. It is the number of particles that delivers a mole of a substance. **Avogadro's number =  $6.022 \times 10^{23}$ .**

**The reason why the value is an accident of nature is that the mole is tied to the gram mass unit.**

The gram is a convenient mass unit because it matches human sizes. If we were a thousand times greater in size ( like Paul Bunyan) we would find it handy to use kilogram amounts. This means the kilogram mole would be convenient. The number of particles handled in a kilogram mole is 1000

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**EW Minium**



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