

## Determining Empirical Formulas

1) 75% carbon, 25% hydrogen	Answer:	$\text{CH}_4$
Work: $\text{C } 75\% \rightarrow 75\text{g C} \left  \frac{1 \text{ mol}}{12.01\text{g C}} = 6.24 \text{ mol} / 6.24 = 1 \right.$ $\text{H } 25\% \rightarrow 25\text{g H} \left  \frac{1 \text{ mol H}}{1.01\text{g H}} = 24.75 \text{ mol} / 6.24 = 4 \right.$ <div style="float: right; border: 1px solid black; padding: 5px;"><math>\text{CH}_4</math></div>		
2) 52.7% potassium, 47.3% chlorine	Answer:	$\text{KCl}$
Work: $\text{K } 52.7\% \rightarrow 52.7\text{g K} \left  \frac{1 \text{ mol K}}{39.10\text{g K}} = 1.34 \text{ mol} / 1.33 = 1 \right.$ $\text{Cl } 47.3\% \rightarrow 47.3\text{g Cl} \left  \frac{1 \text{ mol Cl}}{35.45\text{g Cl}} = 1.33 \text{ mol} / 1.33 = 1 \right.$ <div style="float: right; border: 1px solid black; padding: 5px;"><math>\text{KCl}</math></div>		
3) 22.1% aluminum, 25.4% phosphorous, 52.5 % oxygen	Answer:	$\text{AlPO}_4$
Work: $\text{Al } 22.1\% \rightarrow 22.1\text{g Al} \left  \frac{1 \text{ mol Al}}{26.98\text{g Al}} = 0.819 \text{ mol} / 0.819 = 1 \right.$ $\text{P } 25.4\% \rightarrow 25.4\text{g P} \left  \frac{1 \text{ mol P}}{30.97\text{g P}} = 0.820 \text{ mol} / 0.819 = 1 \right.$ $\text{O } 52.5\% \rightarrow 52.5\text{g O} \left  \frac{1 \text{ mol O}}{16.00\text{g O}} = 3.28 \text{ mol} / 0.819 = 4 \right.$ <div style="float: right; border: 1px solid black; padding: 5px;"><math>\text{AlPO}_4</math></div>		
4) 13% magnesium, 87% bromine	Answer:	$\text{MgBr}_2$
Work: $13\text{g Mg} \left  \frac{1 \text{ mol Mg}}{24.31\text{g Mg}} = 0.54 \text{ mol Mg} / 0.54 = 1 \right.$ $87\text{g Br} \left  \frac{1 \text{ mol Br}}{79.90\text{g Br}} = 1.09 \text{ mol Br} / 0.54 = 2 \right.$ <div style="float: right; border: 1px solid black; padding: 5px;"><math>\text{MgBr}_2</math></div>		
5) 32.4% sodium, 22.5% sulfur, 45.1% oxygen	Answer:	$\text{Na}_2\text{SO}_4$
Work: $32.4\text{g Na} \left  \frac{1 \text{ mol}}{22.99\text{g Na}} = 1.41 \text{ mol Na} / 0.70 = 2 \right.$ $22.5\text{g S} \left  \frac{1 \text{ mol S}}{32.37\text{g S}} = 0.70 \text{ mol S} / 0.70 = 1 \right.$ $45.1\text{g O} \left  \frac{1 \text{ mol O}}{16\text{g O}} = 2.82 \text{ mol O} / 0.70 = 4 \right.$ <div style="float: right; border: 1px solid black; padding: 5px;"><math>\text{Na}_2\text{SO}_4</math></div>		

# Chemistry If8766 Determining Empirical Formulas Answers

**Theodore Rudolph Running**



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