

Matter and Change, pp. 1-4

2. 5. c
3. 5. c
4. 7. c
5. 3. a
6. mass and volume
7. group or family
8. homogeneous
9. products
10. metal
11. solid state
12. chemical properties
13. atom
14. compound
15. nonmetal
16. mixture
17. pure substance
18. basic
19. metalloid
20. reactants: carbon and oxygen; product: carbon dioxide
21. reactant: mercury(II) oxide; products: mercury and oxygen
22. f
23. g
24. c
25. a
26. b
27. i
28. d
29. h
30. e
31. physical change
32. chemical change
33. physical change
34. physical change
35. chemical change
36. physical change
37. In a solid, particles are packed together in relatively fixed positions. The particles vibrate about a fixed point. In a liquid, particles are close together but can flow around one another. In a gas, particles are at great distances from one another, compared to the particles of liquids and solids.
38. The composition of a pure substance is the same throughout and does not vary from sample to sample. Pure substances include elements and compounds. A mixture contains more than one substance and can vary in composition from one sample to another and sometimes within different parts of the same sample.
39. Although energy can be absorbed or released in a

change, it is not destroyed or created. It simply takes a different form.

40. Homogeneous mixtures are uniform in composition; heterogeneous mixtures are not.
41. An element cannot be broken down, but water can be decomposed into hydrogen and oxygen by passing an electrical current through it (electrolysis).

Measurements and Calculations, pp. 5-8

- a 5. c
2. c 6. c
3. c 7. a
4. a 8. a
5. time
6. mass
7. density
8. energy
9. length or distance
10. volume
11. area
12. qualitative
13. quantitative
14. qualitative
15. quantitative
16. $3.00 \times 10^3 \text{ km/s}$
17. three
18. 0.026 g or $2.6 \times 10^{-2} \text{ g}$
19. $2.5 \times 10^{-4} \text{ L}$
20. quantity
21. derived units
22. conversion factor
23. 0.0432 kg or $4.32 \times 10^{-2} \text{ kg}$
24. 5400 mL or $5.4 \times 10^3 \text{ mL}$
25. $300. \text{ K}$
26. $1.05 \times 10^3 \text{ J}$
27. $3.51 \times 10^3 \text{ cm}$
28. A hypothesis is a testable statement that can be used to make predictions and to carry out further experiments. A theory is a broad generalization that explains a body of facts or phenomena.
29. Mass is a measure of the quantity of matter. Weight is a measure of the gravitational pull on matter. Mass does not depend on gravitational attraction.
30. Models are visual, verbal, or mathematical representations. They are used to explain how phenomena occur or how data or events are related.
31. The graph of two quantities that are directly proportional is a straight line through the origin.

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Answers Stoichiometry

**Holt Rinehart & Winston, Holt,
Rinehart and Winston Staff**



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