

Percent Yield Practice

3 pages + 3-page Answer KEY + Periodic Table

Chemistry

Name _____

Date _____

PERCENT YIELD

Show all work and circle or highlight your final answer.

$$\text{percent yield (\%)} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100$$

- 1 10.0 grams of sodium reacts with excess chlorine gas.



The theoretical yield of sodium chloride is 25.0 grams. The actual yield was 21.5 grams. Calculate the percent yield of sodium chloride.

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- 2 Hydrogen gas (H_2) reacts with oxygen gas (O_2) to produce water (H_2O). 5.0 grams of hydrogen reacts with excess oxygen gas.



What is the theoretical yield?

- 3 4.0 grams of carbon reacts with excess oxygen gas.



The theoretical yield of carbon dioxide is 29.3 grams. Calculate the percent yield of carbon dioxide.

- 4 1.0 grams of zinc reacts with excess sulfur.



The theoretical yield of zinc sulfide is 15.5 grams. Calculate the percent yield of zinc sulfide.

- 5 The actual yield was 18.0 grams. Calculate the percent yield.

- 6 Magnesium (Mg) reacts with hydrochloric acid (HCl) to produce magnesium chloride (MgCl_2) and hydrogen gas (H_2). 10.0 grams of magnesium reacts with excess hydrochloric acid.



What is the theoretical yield?

- 7 5.0 grams of phosphorus reacts with excess chlorine gas.



The theoretical yield of phosphorus trichloride is 15.0 grams. Calculate the percent yield of phosphorus trichloride.

- 8 The actual yield was 40.0 grams. Calculate the percent yield.

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- 9 Calcium carbonate (CaCO_3) reacts with hydrochloric acid (HCl) to produce calcium chloride (CaCl_2), carbon dioxide (CO_2), and water (H_2O). 15.0 grams of calcium carbonate reacts with excess hydrochloric acid.



What is the theoretical yield of calcium chloride?

$$\begin{array}{l} 15.0 \text{ g CaCO}_3 \times \frac{1 \text{ mol CaCO}_3}{100.09 \text{ g CaCO}_3} \\ = 0.15 \text{ mol CaCO}_3 \times \frac{1 \text{ mol CaCl}_2}{1 \text{ mol CaCO}_3} \\ = 0.15 \text{ mol CaCl}_2 \end{array}$$

$$\begin{array}{l} \text{Molar Mass CaCl}_2 \\ \text{Ca} = 40.08 \\ \text{Cl} = 35.45 \\ 2 \times 35.45 = 70.90 \\ 40.08 + 70.90 = 110.98 \text{ g} \end{array}$$

$$\begin{array}{l} 0.15 \text{ mol CaCl}_2 \times \frac{1 \text{ mol CaCl}_2}{1 \text{ mol CaCl}_2} \times 110.98 \text{ g CaCl}_2 \\ = 0.15 \times 110.98 \text{ g CaCl}_2 \\ = 16.65 \text{ g CaCl}_2 \end{array}$$

$$\begin{array}{l} \text{Molar Mass CaCl}_2 \\ \text{Ca} = 40.08 \\ \text{Cl} = 35.45 \times 2 = 70.90 \\ 40.08 + 70.90 = 110.98 \text{ g} \end{array}$$

- 10 The actual yield was 15.5 grams. Calculate the percent yield of calcium chloride.

$$\text{percent yield (\%)} = \frac{15.5 \text{ g}}{16.65 \text{ g}} \times 100 = 92.8\%$$

- 11 Potassium (K) reacts with iodine (I_2) to produce potassium iodide (KI). 1.0 grams of potassium reacts with excess iodine.



What is the theoretical yield?

$$\begin{array}{l} 1.0 \text{ g K} \times \frac{1 \text{ mol K}}{39.10 \text{ g K}} \\ = 0.0256 \text{ mol K} \times \frac{2 \text{ mol KI}}{2 \text{ mol K}} \\ = 0.0256 \text{ mol KI} \end{array}$$

$$\begin{array}{l} \text{Molar Mass KI} \\ \text{K} = 39.10 \\ \text{I} = 126.90 \\ 39.10 + 126.90 = 166.00 \text{ g} \end{array}$$

$$\begin{array}{l} 0.0256 \text{ mol KI} \times \frac{1 \text{ mol KI}}{1 \text{ mol KI}} \times 166.00 \text{ g KI} \\ = 0.0256 \times 166.00 \text{ g KI} \\ = 4.25 \text{ g KI} \end{array}$$

YIELD

ANSWER KEY INCLUDED

Chemistry Matter And Change Percent Yield Section 1study Guide Answer

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The Complete Chemistry Study Guide and Note Cards and MCAT Konstantinos Papadopoulos, 2012-07-06 **Chemistry:**
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Chemistry for Today Seager/Slabaugh, 2004 The fifth edition of the Study Guide and Student Solutions Manual has been
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review of each section of the chapter and a section entitled Tying It All Together with a Laboratory Application **Chemical**
Engineer ,1909 Chemical Age ,1910 **Pharmaceutical Journal** ,1900 **Industries** ,1887 Brewers' guardian
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Grain Trade ,1920 **The Chemical Age** ,1922 *Liverpool and Manchester Photographic Journal* William Crookes, T.A.
Malone, George Shadbolt, J. Traill Taylor, William Blanchard Bolton, Thomas Bedding, 1897 **Merck's Report** Theodore
Weicker, 1897 *New Zealand Farmer Weekly* ,1923 **Leslie's Illustrated Weekly Newspaper** ,1917 **The Journal**
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