

### balance nuclear reaction equations

aluminum-26 undergoes positron emission  
potassium-40 undergoes beta decay (electron)  
uranium-238 undergoes alpha decay



### electron configurations

Write the electron configuration for C.

Write the electron configuration for Cr.

Write the shorthand notation for Bi.



### NOMENCLATURE [naming compounds, writing formulas]

Write the formula for each of the following compounds:

- a. Sulfur hexafluoride  $\text{SF}_6$
- b. Lithium nitride  $\text{Li}_3\text{N}$
- c. Chromium (III) carbonate  $\text{Cr}_2(\text{CO}_3)_3$
- d. Tin (II) chloride  $\text{SnCl}_2$
- e. Ammonium acetate  $\text{NH}_4\text{C}_2\text{H}_3\text{O}_2$  (or  $\text{CH}_3\text{COONH}_4$ )
- f. Mercury (I) chloride  $\text{Hg}_2\text{Cl}_2$
- g. Potassium bromate  $\text{KBrO}_3$
- h. Hydrosulfuric acid  $\text{H}_2\text{S(aq)}$
- i. Chloric acid  $\text{HClO}_3$
- j. Sulfurous acid  $\text{H}_2\text{SO}_3$

Name each of the following compounds.

- a.  $\text{CuSO}_4$  Copper (II) sulfate
- b.  $\text{AlF}_3$  aluminum fluoride
- c.  $\text{HI}$  hydrogen iodide
- d.  $\text{NO}$  nitrogen monoxide
- e.  $\text{H}_2\text{Se(aq)}$  hydroselenic acid
- f.  $\text{HNO}_3$  nitric acid
- g.  $\text{HNO}_2$  nitrous acid
- h.  $\text{NaHSO}_4$  sodium hydrogen sulfate

### Half-life

Tritium ( $\text{H-3}$ ) is a radioactive isotope of hydrogen with a half-life of 12.3 years. How long would it take for a 40.0 g sample to decay down to 1.25 g?

$$40 \rightarrow 20 \rightarrow 10 \rightarrow 5 \rightarrow 2.5 \rightarrow 1.25$$

1 2 3 4 5 61.5 yr

Fe-61 has a half-life of 6.00 min. Of a 100.0 mg sample, how much will remain after 18.0 min?

$$100 \rightarrow 50 \rightarrow 25 \rightarrow 12.5 \text{ mg}$$

1 2 3 12.5 mg

3 half lives

### Average atomic mass

The element copper has naturally occurring isotopes with mass numbers of 63 and 65.

The relative abundance and atomic masses are 69.2% for a mass of 63amu and 30.8% for a mass of 65amu. Calculate the average atomic mass of copper.

$$\begin{array}{rcl} 63 & \times & .692 = 43.596 \rightarrow 44 \\ 65 & \times & .308 = 20.02 \rightarrow 20. \\ & + & \\ & \hline & & 63.62 \rightarrow 64 \end{array}$$

# Chemistry Of Life Study Guide Answer Key

**Susan McMurry, John McMurry**



## Chemistry Of Life Study Guide Answer Key:

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**The Publishers' Trade List Annual**, 1904      **Study Guide for Life on Earth** Carl S. Pike, 1973      *Life* William K. Purves, 2001. Authoritative, thorough and engaging. *Life: The Science of Biology* achieves an optimal balance of scholarship and teachability, never losing sight of either the science or the student. The first introductory text to present biological concepts through the research that revealed them. *Life* covers the full range of topics with an integrated experimental focus that flows naturally from the narrative. This approach helps to bring the drama of classic and cutting edge research to the classroom but always in the context of reinforcing core ideas and the innovative scientific thinking behind them. Students will experience biology not just as a litany of facts or a highlight reel of experiments but as a rich, coherent discipline.

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**and Biological Chemistry** Susan McMurry, 2006-07 The Study Guide and Full Solutions Manual explain in detail how the answers in text and end of chapter problems are obtained They also contain chapter summaries study hints and self tests for each chapter *Study guide to accompany Drew H. Wolfe: General, organic and biological chemistry* Daniel C. Pantaleo, 1986 **U.S. Naval Training Bulletin** , 1947-10 El-Hi Textbooks in Print , 1984 **Study Guide for Organic Chemistry** Seyhan N. Ege, 1999 *El-Hi Textbooks and Serials in Print* , 2006 **Practicing Biology** Jean Heitz, Jane B. Reece, 2004 Table of contents continued How are water and good transported in plants What do you need to consider in order to grow plants in space or anywhere else for that matter How can plant reproduction be modified using biotechnology How do gravity and light affect plant growth responses How does an organism s structure help it maintain homeostasis How are form and function related in the digestive system How is mammalian heart structure related to function How do we breathe and why do we breathe How does the immune system keep the body free of pathogens What is nitrogenous waste and how is it removed from the body How do hormones regulate cell functions How does the production of male and female gametes differ in humans What common events occur in the early development of animals How do neurons function to transmit information What would happen if you modified a particular aspect of neuron function How does sarcomere structure affect muscle function What would happen if you modified particular aspects of muscle function What factors determine climate What determines behavior What methods can you use to determine population density and distribution What models can you use to calculate how quickly a population can grow What do you need to consider when analyzing communities of organisms What limits do available solar radiation and nutrients place on carrying capacities What factors can affect the survival of a species or community The activities of this workbook focus on key ideas principles and concepts that are basic to understanding biology The overall organization follows that of Campbell Reece Biology 7th edition p vii **Organic and Biological Chemistry** Karen Timberlake, 2001-11 Keyed to the learning goals in the Organic and Biological Chemistry text this guide is designed to promote active learning through a variety of exercises with answers and mastery exams The guide also contains complete solutions to odd numbered problems Subject Guide to Books in Print , 1993 Biological Sciences Curriculum Study Journal Biological Sciences Curriculum Study, 1978

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