

Reactions of Amines and Epoxides

Preparation of Amines

Amination of alcohols: Polyalkylation is a serious problem for indirect synthesis of amines



A Summary of Ring-Opening Reactions of Epoxides with Strong and Weak Nucleophiles



Bond Types, Strength, Length, Polarity and Dipole Moment

Two types of covalent bond

There are two types of covalent bond: ionic bond and covalent bond

- Ionic bonds are formed by the transfer of electrons from one element to another (ionization potential) and electron affinity (electron affinity).



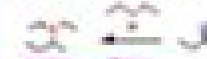
- Covalent bonds are formed by sharing of electrons between two atoms.



The Relative Electronegativity of Atoms



Bond and Bonding in a Molecule



Bonding in a Molecule



There are two types of bond

Definition of ionic bond



Sigma and pi bonds

Covalent bonds are classified into sigma and pi bonds

- Sigma bonds are formed by head-on overlap of orbitals.
- pi bonds are formed by side-on overlap of orbitals.
- One single bond is a sigma bond.
- One double bond consists of one sigma and one pi bond.



pi bonds are stronger than sigma bonds

Polar and Nonpolar Covalent Bonds

Depending on the difference in the electronegativity values, covalent bonds are divided into polar and nonpolar.



- If the difference in electronegativity is less than 0.5, the electrons are shared equally (nonpolar bond). For example, C-C and C-H bonds.

- If the difference in electronegativity is between 0.5 and 1.7, the electrons are shared unequally (polar bond). For example, C-O and C-N bonds.

- A difference of 1.7 or higher is ionic bond (the electrons are not shared).

Dipole Moment

The polarity of a bond is described by the dipole moment (μ).



- The dipole moment is a vector pointing from the less electronegative atom to the more electronegative atom.
- The more electronegative atom has a partial negative charge (δ^-) and the less electronegative atom has a partial positive charge (δ^+).
- δ^- (delta minus) indicates an electron deficient atom.
- δ^+ (delta plus) indicates an electron rich atom.

Molecular Dipole

To determine the molecular polarity, the vector sum of the dipole moments

The polarity of the molecule is determined by the vector sum of the dipole moments. A good example is carbon dioxide.



Carbon dioxide is a linear molecule and the dipole moment is zero.

Bond Length and Bond Strength

The length of the bond is measured in Å.

Bond	H-H	H-Cl	H-Br	H-I
Length	0.074 Å	0.127 Å	0.141 Å	0.154 Å
Strength	436 kJ/mol	431 kJ/mol	366 kJ/mol	298 kJ/mol

Double bonds are stronger than single bonds and are shorter.

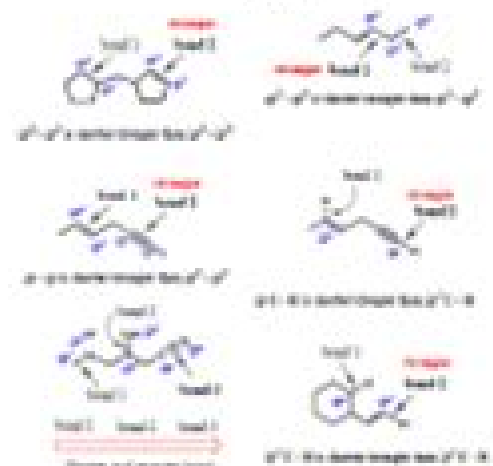


The length of the bond is measured in Å.

Average values of C-C and C-H bonds depending on hybridization			
C-C Bond	sp^3-sp^3	sp^2-sp^2	$sp-sp$
Hybridization	sp^3-sp^3	sp^2-sp^2	$sp-sp$
Length	1.54 Å	1.34 Å	1.20 Å
Strength	376 kJ/mol	611 kJ/mol	974 kJ/mol
Hybridization	sp^3-sp^2	sp^2-sp	$sp-sp$
Length	1.50 Å	1.43 Å	1.20 Å
Strength	418 kJ/mol	602 kJ/mol	1013 kJ/mol

Higher s character means shorter and stronger bond

Example



Chemistry Of Organic Molecules Study Guide Answers

**T. W. Graham Solomons, Craig B.
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