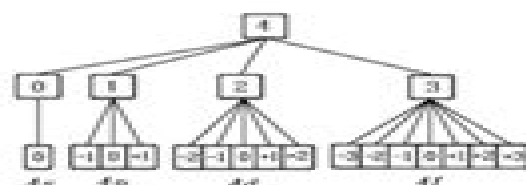


Chapter 1

Exercises

- 1.1 (a) Surface where electron probability = 0.
 (b) No two electrons in an atom can have the same 4 quantum numbers.
 (c) Attraction into a magnetic field by an unpaired electron.

1.3



- 1.5 $5p$.
 1.7 Size of an orbital.
 1.9 With parallel spins there is zero probability that the electrons will occupy the same volume of space.
 1.11 (a) $[\text{Ne}]3s^1$, (b) $[\text{Ar}]4s^2 3d^8$, (c) $[\text{Ar}]4s^1 3d^1$.
 1.13 (a) $[\text{Ar}]$, (b) $[\text{Ar}]$, (c) $[\text{Ar}] 3d^9$.
 1.15 $1+$ and $3+$. Configuration of $[\text{Xe}]6s^2 4f^{14} 5d^1 6p^1$.
 1.17 $1+$. Configuration of $[\text{Kr}]5s^1 4d^1$.
 1.19



- 1.21 E 113: $[\text{Rn}]7s^2 5f^{14} 6d^{10} 7p^1$.
 E 113 as $+1$ ion: $[\text{Rn}]7s^2 5f^{14} 6d^{10}$.
 E 113 as $+3$ ion: $[\text{Rn}]5f^{14} 6d^{10}$.

Beyond the Basics

- 1.23 9, 5, 121.
 1.25 There are seven f orbitals. There are two separate ways of depicting them and designating them: the general set and the cubic set.
 1.27 Hydrogen heads the alkali metal group. Helium heads the alkaline earth metal group.

Chapter 2

Exercises

- 2.1 (a) Lanthanum through lutetium.
 (b) Apparent radius of an atom in non-bonded contact with another.
 (c) Actual nuclear charge experienced by an electron.
 2.3 Argon did not fit into any of the then-known groups. Because the table was based on measured atomic mass, argon should have been placed between potassium and calcium.
 2.5 The long form correctly depicts the order of elements, but the table becomes very elongated.
 2.7 The $-m$ ending indicates a metal. The ending $-on$ has been used for non-metals.
 2.9 With nuclei up to 26 protons, nuclear fusion is an exothermic process.
 2.11 (a) Lead (b) technetium (c) bromine.
 2.13 Sodium, because it has an odd number of protons.
 2.15 50.
 2.17 (a) Several nonmetals have metallic luster.
 (b) Diamond has the highest thermal conductivity of all substances.
 (c) Graphite is a good electrical conductor in two dimensions.
 2.19 Potassium. As the effective atomic charge on the outermost electrons increases across a period.
 2.21 The effective nuclear charge on the $4p$ electrons will be increased.

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