

Задача 11
Слабые и сред. кислоты

1) Дано:

$$C_{HNO_2} = 0,01 \text{ м}$$

$$K_a = 7,9 \cdot 10^{-4}$$

$$pH = ?$$

$$pH = -\lg [H^+] = -\lg 2,8 \cdot 10^{-5} = 4,55$$

Решение:

Для слабой кислоты:

$$[H^+] = \sqrt{K_a \cdot C_{HNO_2}}$$

$$= \sqrt{7,9 \cdot 10^{-4} \cdot 10^{-2}} = 2,8 \cdot 10^{-5} \text{ моль/л}$$

$$= 4,55$$

2) Дано:

$$K_{Pr} = 6,3 \cdot 10^{-5}$$

$$pH = ?$$

ξ - гидролизная степень

$$\xi = \sqrt{\frac{K_{Pr}}{C}} = \sqrt{\frac{6,3 \cdot 10^{-5}}{4}} = 1,25 \cdot 10^{-2}$$

$$[OH^-] = \frac{\xi}{2} = \frac{1,25 \cdot 10^{-2}}{2} = 6,25 \cdot 10^{-3} \text{ моль/л}$$

$$[H^+] = \frac{10^{-14}}{[OH^-]} = \frac{10^{-14}}{6,25 \cdot 10^{-3}} = 1,6 \cdot 10^{-12}$$

$$pH = -\lg [H^+] = -\lg 1,6 \cdot 10^{-12} = 11,8$$

Решение:

$$C_{NaOH} = C_{H_2CO_3} = 2,5 \cdot 10^{-2}$$

$$K_{Pr} = [CO_3^{2-}][OH^-]$$

$$= 5 \cdot (2,5)^2 = 31,25$$

3) Дано:

$$V_{HCl} = 10 \text{ мл}$$

$$C_{HCl} = 0,1 \text{ м}$$

$$V_{NaClO_2} = 10 \text{ мл}$$

$$C_{NaClO_2} = 0,1 \text{ м}$$

$$K_{Pr} = 4,6 \cdot 10^{-5}$$

$$PrCl_2 \rightleftharpoons Pr^{2+} + 2Cl^-$$

Решение:

полн. соль

$$C_{HCl} = \frac{0,1 \cdot 10}{20} = 0,05 \text{ моль/л}$$

$$C_{NaClO_2} = \frac{0,1 \cdot 10}{20} = 0,05 \text{ моль/л}$$

полн гидролизующая соль

$$C_{Cl^-} = C_{HCl} = 0,05 \text{ моль/л}$$

$$C_{Pr^{2+}} = C_{NaClO_2} = 0,05 \text{ моль/л}$$

$$K_{Pr} = [Pr^{2+}][Cl^-]^2 = 4,6 \cdot 10^{-5} = 10^{-4,34}$$

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