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Student Exploration: Boyle's Law and Charles's Law

Vocabulary: absolute zero, Boyle's law, Charles's law, Gay-Lussac's law, Kelvin scale, pressure, obs

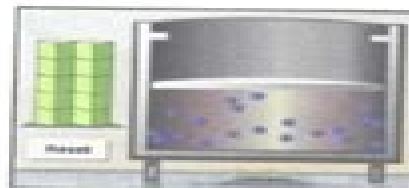
Prior Knowledge Question (Do this BEFORE using the Gizmo.)

A small helium tank measures about two feet (60 cm) high. Yet it can fill over 50 balloons! How can such a small tank contain enough helium to fill so many balloons?

A helium tank has a lot of helium molecules inside but you do not need a lot of molecules to fill up a balloon. Therefore, a small tank can fill up a lot of balloons.

Gizmo Warm-up

The Boyle's Law and Charles's Law Gizmo shows a container of gas. Inside, small purple spheres represent gas molecules.



1. Observe the particles. Are they all moving at the same

speed? No some particles are faster than others

2. How do the particles interact with the walls and lid of the container? The particles

bounce off the walls and lid of the container and bump with other particles.

These interactions contribute to the **pressure** on the walls of the container. Pressure is defined as force per unit area. The SI units of pressure are newtons per square meter (N/m^2), or pascals (Pa).

3. Slowly drag the temperature (T) slider back and forth. (Note: In this Gizmo, the Kelvin scale is used to measure temperature. On the Kelvin scale, 0 degrees is **absolute zero**, the coldest possible temperature. Absolute zero is equal to -273.15°C or -459.67°F .)

- A. How does the change in temperature affect the speed of the molecules? if the

temperature goes up the particles move faster and if it goes down the particles are more compact and move slower

- B. How does the change in temperature affect the volume of the container? if the temperature

goes up the particles will have more space to roam around and the container will increase in volume.

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Ian Pickup



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