

### balance nuclear reaction equations

aluminum-26 undergoes positron emission  
potassium-40 undergoes beta decay (electron)  
uranium-238 undergoes alpha decay



### electron configurations

Write the electron configuration for C.

Write the electron configuration for Cr.

Write the shorthand notation for Bi.



### NOMENCLATURE [naming compounds, writing formulas]

Write the formula for each of the following compounds:

- a. Sulfur hexafluoride  $\text{SF}_6$
- b. Lithium nitride  $\text{Li}_3\text{N}$
- c. Chromium (III) carbonate  $\text{Cr}_2(\text{CO}_3)_3$
- d. Tin (II) chloride  $\text{SnCl}_2$
- e. Ammonium acetate  $\text{NH}_4\text{C}_2\text{H}_3\text{O}_2$  (or  $\text{CH}_3\text{COONH}_4$ )
- f. Mercury (I) chloride  $\text{Hg}_2\text{Cl}_2$
- g. Potassium bromate  $\text{KBrO}_3$
- h. Hydrosulfuric acid  $\text{H}_2\text{S(aq)}$
- i. Chloric acid  $\text{HClO}_3$
- j. Sulfurous acid  $\text{H}_2\text{SO}_3$

Name each of the following compounds.

- a.  $\text{CuSO}_4$  Copper (II) sulfate
- b.  $\text{AlF}_3$  aluminum fluoride
- c.  $\text{HI}$  hydrogen iodide
- d.  $\text{NO}$  nitrogen monoxide
- e.  $\text{H}_2\text{Se(aq)}$  hydroselenic acid
- f.  $\text{HNO}_3$  nitric acid
- g.  $\text{HNO}_2$  nitrous acid
- h.  $\text{NaHSO}_4$  sodium hydrogen sulfate

### Half-life

Tritium ( $\text{H-3}$ ) is a radioactive isotope of hydrogen with a half-life of 12.3 years. How long would it take for a 40.0 g sample to decay down to 1.25 g?

$$40 \xrightarrow{1} 20 \xrightarrow{2} 10 \xrightarrow{3} 5 \xrightarrow{4} 2.5 \xrightarrow{5} 1.25$$

Fe-61 has a half-life of 6.00 min. Of a 100.0 mg sample, how much will remain after 18.0 min?

$$3 \text{ half lives} \quad 100 \xrightarrow{1} 50 \xrightarrow{2} 25 \xrightarrow{3} 12.5 \text{ mg}$$

### Average atomic mass

The element copper has naturally occurring isotopes with mass numbers of 63 and 65.

The relative abundance and atomic masses are 69.2% for a mass of 63amu and 30.8% for a mass of 65amu. Calculate the average atomic mass of copper.

$$\begin{array}{rcl} 63 \times .692 & = & 43.596 \rightarrow 44 \\ 65 \times .308 & = & 20.02 \rightarrow 20. \\ \hline & + & \\ & 63.62 & \rightarrow 64 \end{array}$$

# Chemistry Answers Guided Practice Problem

**Anita Woolfolk, Kay Margetts**



## **Chemistry Answers Guided Practice Problem:**

*Survival Guide to General Chemistry* Patrick E. McMahon, Rosemary McMahon, Bohdan Khomtchouk, 2019-02-13 This work evolved over thirty combined years of teaching general chemistry to a variety of student demographics. The focus is not to recap or review the theoretical concepts well described in the available texts. Instead, the topics and descriptions in this book make available specific detailed step by step methods and procedures for solving the major types of problems in general chemistry. Explanations, instructional process sequences, solved examples, and completely solved practice problems are greatly expanded, containing significantly more detail than can usually be devoted to in a comprehensive text. Many chapters also provide alternative viewpoints as an aid to understanding. Key Features: The authors have included every major topic in the first semester of general chemistry and most major topics from the second semester. Each is written in a specific and detailed step by step process for problem solving, whether mathematical or conceptual. Each topic has greatly expanded examples and solved practice problems containing significantly more detail than found in comprehensive texts. Includes a chapter designed to eliminate confusion concerning acid base reactions, which often persists through working with acid base equilibrium. Many chapters provide alternative viewpoints as an aid to understanding. This book addresses a very real need for a large number of incoming freshmen in STEM fields.

***Survival Guide to Organic Chemistry*** Patrick E. McMahon, Bohdan B. Khomtchouk, Claes Wahlestedt, 2016-12-19 Reviews key general chemistry concepts and techniques adapted for application to important organic principles. Provides practical guidance to help students make the notoriously well known and arduous transition from general chemistry to organic chemistry. Explains organic concepts and reaction mechanisms, generally expanding the focus on how to understand each step from a more intuitive viewpoint. Covers concepts that need further explanation as well as those that summarize and emphasize key ideas or skills necessary in this field. An added bonus is help with organizing principles to make sense of a wide range of similar reactions and mechanisms. Implements a user friendly process to achieve the end result of problem solving. Covers organic chemistry I and II concepts at the level and depth of a standard ACS organic chemistry curriculum. Features practice problems and solutions to help master the material, including an extensive and comprehensive bank of practice exams with solutions.

*Guide to Mining Examinations with Answers to Examination Questions* Henry Davies, 1894 *Problem Solving Guide and Workbook for Introductory Chemistry* by Steve Russo, Mike Silver, Sandra Yancy McGuire, Steve Russo, Mike Silver, 2002 Provides over 175 worked examples and more than 500 practice problems and quiz questions to help students develop and practice their problem solving skills. *Study Guide with Selected Solutions* Karen Timberlake, 2002-08-06 Written by the author, the Study Guide is keyed to the learning goals in the text and designed to promote active learning through a variety of exercises with answers and mastery exams. Also contains complete solutions to odd numbered problems. *Study Guide and Partial Solutions Manual, Fundamentals of General, Organic, and Biological Chemistry* Susan McMurry, 1996 This internationally acclaimed detective series is just the

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Study Guide and Full Solutions Manual Susan McMurry,John McMurry,2002-08 Contains a brief overview of every chapter review of skills self tests and the answers and detailed solutions to all end of chapter problems in the textbook

**Student's Guide to Chemistry; a Modern Introduction** David Brooks,1974      **Student Study Guide**  
**Fundamentals of Chemistry by David E. Goldberg** Margaret G. Kimble,1994      **Instructor's Guide for Problem Solving and Comprehension** Arthur Whimbey,Jack Lochhead,1982      Chemistry John S. Phillips,Cheryl Wistrom,2000

Introduction to Chemistry Martha J. Gilleland,Rebecca Williams,1986      **Student Study Guide to Accompany Chemistry** Kenneth J. Hughes,Paul B. Kelter,2002-05 Reviews essential points from the text presents problems and hints to assist the student in solving them Solutions to selected problems in the end of chapter problem sets are included By Kenneth Hughes of the University of Wisconsin Madison      Addison-Wesley Chemistry Antony C. Wilbraham,Dennis D.

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