

CHEMISTRY COMES ALIVE

Matter – anything with mass and occupies space

- weight – pull of gravity on mass
- solid – with definite shape and volume
- liquid – with definite volume but changeable shape
- gas – with changeable volume and shape

Energy - capacity to do work or put matter into motion

- 2 types – kinetic energy – energy in action
- potential energy – stored (inactive) energy
- energy can be transferred from potential to kinetic

Forms of energy used by the body:

1. Chemical energy – stored in the chemical bonds
 - ATP (energy of the body)
2. Electrical energy - result from movement of charged particles
3. Mechanical energy – energy that directly cause the movement of matter
4. Radiant or electromagnetic energy - energy that travel in waves.
 - light, infrared waves, radiowaves, ultraviolet waves, X-rays

Energy - is neither created nor destroyed

- can be converted from one form to another form,
- some energy is unused in the conversion (lighting a bulb)

Composition of Matters: All matters are made up of elements

Elements: - unique substances that cannot be broken into simpler form by ordinary chemical means

- 2 properties of elements – physical – anything we can sense or measure
- chemical – how atoms interact (bond) with other atoms

Atoms -unique building blocks of elements

- give each element its chemical and physical properties
- smallest particles of the element that retain the properties of the element
- atomic symbols – designated for each element.
 - usually represented by the first letter or the first 2 letter of the name of the element in English, Latin mainly and rarely in other languages
- 118 elements – 92 are naturally occurring
- 26 are produced from natural element by particle accelerator or nuclear reactor.

Chemistry Comes Alive Study Guide Answer Key

Arthur James Wells



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Student Guide For Living Chemistry Barbara Ucko, 2012-12-02 Student Guide for Living Chemistry is a 23 chapter textbook guide that allows students to study and review on their own and test their understanding to help them prepare for examinations Every chapter begins with a list of objectives stating exactly the skills to develop in a particular unit Each objective corresponds to a section in the textbook Living Chemistry Three kinds of questions are provided for each objective to check the student's understanding namely short answer Study Questions multiple choice and fill in The answers for all questions are provided at the end of the chapter The opening chapters cover the SI units composition of matter chemical bonding compounds chemical change gases respiration and water The subsequent chapters deal with solutions acids bases salts nuclear and organic chemistry oxygen derivatives and hydrocarbons polymers and other organic derivatives This textbook also explores the chemistry of carbohydrates lipids proteins enzymes and energy and carbohydrate metabolism The remaining chapters discuss the chemistry of vitamins hormones body fluid drugs and poisons Undergraduate chemistry students will find this book invaluable

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