

Study Guide – Chapter 13: Solutions

Test Date: _____

Chapter Four:

- Making solutions, Molarity, & Dilutions
 - Molarity Calculations
 - Determine moles, mass, volume, or molarity depending on the info given in the problem (remember that $M = \# \text{ moles solute} / 1 \text{ L solution}$)
 - Dilution – when you take a *more concentrated* solution (like a stock solution) and make a *less concentrated* solution
 - $M_1V_1 = M_2V_2$
 - Making Solutions (from solids and dilutions)
 - Solids: Need to know what volume & concentration you are making → this way you can calculate moles, and thus mass, of the solid you will need.
 - Dilutions: Use the dilution equation to determine what volume of concentrated you will need to dilute down and add then water to desired volume
 - Always use a volumetric flask for solutions → flask can measure 1 volume only

Chapter Thirteen:

- The Solution Process
 - Terms to know: solution, solute, solvent (the solute dissolves into the solvent)
 - ~~Energy Changes of solution process~~
 - ~~three steps → separation of solute particles, separation of solvent particles, formation of solute-solvent interactions (is each step endo or exothermic?)~~
 - ~~When is an overall solution endothermic? exothermic?~~
 - When a solute is being dissolved in a solvent...
 - Dissolution vs. Crystallization – two opposite processes that occur in a solution in dynamic equilibrium (when the rate of dissolution equals the rate of crystallization no more solute can be dissolved in the solution)
 - Solubility Curves & Saturated/Unsaturated/Supersaturated solutions
 - The line represents a SATURATED solution dissolved (typically) in 100 g of water (above line = supersaturated; below line = unsaturated)
- Factors that Affect Solubility of a Solute
 - Nature of the solute and solvent ("Like dissolves Like")
 - polar solvents dissolve polar (and ionic) solutes
 - nonpolar solvents dissolve nonpolar solutes
 - be able to determine what solvent a substance will be more likely do dissolve in based on polarity (like in the review sheet)
 - Terms to know: miscible vs. immiscible
 - Solubility for gases:
 - What increases solubility of a gas? → larger molar mass, increased pressure, decreased (colder) temperatures
 - Solubility for solids:
 - What increases solubility of a solid? → stirring/agitation, increased (warmer) temperatures, decreased surface area (smaller particle size), pressure has NO affect on solids

Chemistry Concepts And Applications Study Guide

Chapter 13 Answers

Kenneth W. Raymond



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