

Experiment 4: Chemical Equilibrium - Testing LeChâtelier's Principle Report Sheet

Part A. Equilibrium State and Changes in Reactant or Product Concentration.

Observations	Copper Sulfate	Nickel Chloride
Initial Color	blue	green
Observation After Adding Concentrated Ammonia	color changes to a very dark royal blue	color changed to a dark blue / purple
Did any solid form?	no	no
Observations After adding 1 M Hydrochloric Acid	a precipitate formed and the color turned a lighter yellowish color	a precipitate formed and the color is a pale green

Account for the observation by writing the balance equation for each reaction.

1. Addition of 15 M ammonia to



2. Based on the chemical reaction of ammonia with water, what could be the nature of the solid momentarily formed while adding 15 M ammonia to each solution?



3. Addition of 1M Hydrochloric acid



Chemistry Exploring Equilibrium Post Lab Answers

J Ma



Chemistry Exploring Equilibrium Post Lab Answers:

Exploring General Chemistry in the Laboratory Colleen F. Craig, Kim N. Gunnerson, 2017-02-01 This laboratory manual is intended for a two semester general chemistry course The procedures are written with the goal of simplifying a complicated and often challenging subject for students by applying concepts to everyday life This lab manual covers topics such as composition of compounds reactivity stoichiometry limiting reactants gas laws calorimetry periodic trends molecular structure spectroscopy kinetics equilibria thermodynamics electrochemistry intermolecular forces solutions and coordination complexes By the end of this course you should have a solid understanding of the basic concepts of chemistry which will give you confidence as you embark on your career in science

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