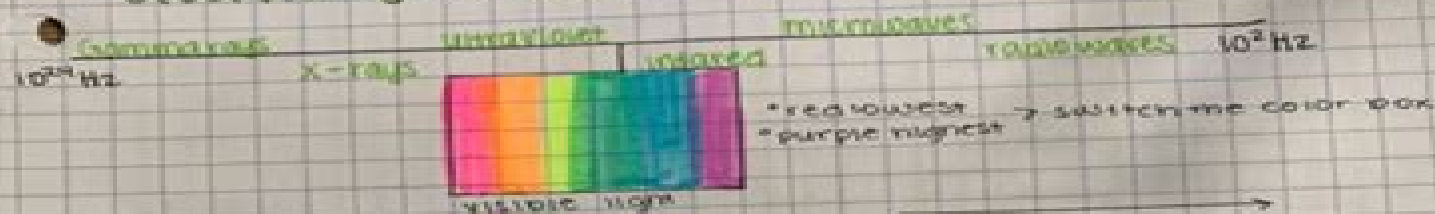


unit 3 - quiz 1

Electromagnetic Radiation



• increase in energy

• frequency increases

• shorter wavelength

• wave length: distance from the top of one wave to the next

• frequency: # of waves that pass a given point per second

• speed of light 3.0×10^8 m/s

• decrease in energy

• frequency decreases

• longer wavelength

Periodic Table Groups

Alkali Metals: Group 1, 1 valence e⁻

Alkaline Metals: Group 2, 2 valence e⁻

Transition Metals: Group 3-10, valence e⁻ varies

Halogens: Group 17, 1 valence e⁻

• as you go left to right on periodic table it goes from metals to non-metals

• majority of elements are metal, arranged by atomic number

• metalloids have metallic and non-metallic properties

• Mendeleev's periodic table

• periodic law: when elements listed in order of atomic # and end up in groups with similar properties

Electron Configuration

• sublevels s, p, d, f, energy levels 1-7

- s one orbital, max 2e⁻

- p three orbital, max 6e⁻, min 2e⁻

- d five orbital, max 10e⁻, min 2e⁻

- f seven orbital, max 14e⁻, min 2e⁻

energy level

of electron in the sublevel

2p⁴
sublevel

Ex. Iodine: $1s^2 2s^2 2p^6 3s^2 3p^4 4s^2 4p^6 5s^2 5p^5$

Noble Gas Configuration

• Find noble gas before element (row above) (Group 18)

• place symbol in brackets

• write remaining electron configuration

Ex. krypton: $[Kr] 4s^2 4p^4 5d^0 5p^0$

dot configuration

• based on valence e⁻, orbital with no electrons is empty

• each single dot = one valence electron

• K • As • Se • S • H •

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