

## SECTION-1 (SINGLE ANSWER TYPE QUESTIONS)

31. Number of atoms in 560 g of Fe (atomic mass  $56 \text{ g mol}^{-1}$ ) is  
 i) twice that of 70 g N  
 ii) half that of 20 g H  
 iii) half that of 40 g He  
 a) i, ii                      b) ii, iii                      c) i, ii, iii                      d) only iii
32. How many moles of magnesium phosphate,  $\text{Mg}_3(\text{PO}_4)_2$  will contain 2.5 mol of oxygen atoms?  
 a) 3.2                      b)  $3.125 \times 10^{-1}$                       c)  $3.125 \times 10^{-2}$                       d) 0.32
33. Experimentally it was found that a metal oxide has formula  $\text{M}_{0.95}\text{O}$ . Metal M, is present as  $\text{M}^{2+}$  and  $\text{M}^{3+}$  in its oxide. Fraction of the metal which exists as  $\text{M}^{3+}$  would be  
 a) 4.08 %                      b) 6.05 %                      c) 95.92 %                      d) 93.95 %
34. The percentage composition of carbon by mole in methane is  
 a) 75 %                      b) 80 %                      c) 25 %                      d) 20 %
35. A 5.2 molal aqueous solution of methyl alcohol  $\text{CH}_3\text{OH}$  is supplied. What is the mole fraction of methyl alcohol in the solution?  
 a) 0.05                      b) 0.090                      c) 0.052                      d) 0.2
36. A solution of sodium sulfate contains 92 g of  $\text{Na}^+$  ions per kilogram of water. The molality of  $\text{Na}^+$  ions in that solution in  $\text{mol kg}^{-1}$  is  
 a) 12                      b) 4                      c) 8                      d) 16
37. 8 g of NaOH is dissolved in 18 g of  $\text{H}_2\text{O}$ . Mole fraction of NaOH in solution and molality (in  $\text{mol kg}^{-1}$ ) of the solution respectively are  
 a) 0.2, 22.20                      b) 0.2, 11.11                      c) 0.167, 11.11                      d) 0.167, 22.20
38. The amount of oxygen in 3.6 mol of water is  
 a) 64.8 g                      b) 57.6 g                      c) 28.8 g                      d) 18.4 g

39. 25 mL of a solution of barium hydroxide on titration with 0.1 molar solution of hydrochloric acid gave a titre value of 35 mL. The molarity of barium hydroxide solution was  
 a) 0.07                      b) 0.14                      c) 0.28                      d) 0.35
40. The volume of 0.1 N dibasic acid sufficient to neutralize 1 g of a base that furnishes  $0.04 \text{ mol}^{-1} \text{OH}^-$  in aqueous solution is  
 a) 400 mL                      b) 600 mL                      c) 200 mL                      d) 800 mL
41. A gaseous compound of nitrogen and hydrogen contains 12.5% (by mass) of hydrogen. The density of the compound relative to hydrogen is 16. The molecular formula of the compound is  
 a)  $\text{NH}_2$                       b)  $\text{N}_3\text{H}$                       c)  $\text{NH}_3$                       d)  $\text{N}_2\text{H}_4$
42. The amount of sugar ( $\text{C}_{12}\text{H}_{22}\text{O}_{11}$ ) required to prepare 2 L of its 0.1 M aqueous solution is  
 a) 136.8 g                      b) 17.1 g                      c) 68.4 g                      d) 34.2 g
43. The molarity of a solution obtained by mixing 750 mL of 0.5 (M) HCl with 250 mL of 2 (M) HCl will be  
 a) 1.00 M                      b) 1.75 M                      c) 0.975 M                      d) 0.875 M
44. For a reaction,  $\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \rightarrow 2\text{NH}_3(\text{g})$   
 Identify dihydrogen ( $\text{H}_2$ ) as a limiting reagent in the following reaction mixtures.  
 a) 56 g of  $\text{N}_2$  + 10 g of  $\text{H}_2$                       b) 35 g of  $\text{N}_2$  + 8 g of  $\text{H}_2$   
 c) 28 g of  $\text{N}_2$  + 6 g of  $\text{H}_2$                       d) 14 g of  $\text{N}_2$  + 4 g of  $\text{H}_2$
45. For the following reaction, the mass of water produced from 445 g of  $\text{C}_{27}\text{H}_{110}\text{O}_6$  is  
 $2\text{C}_{27}\text{H}_{110}\text{O}_6(\text{s}) + 163\text{O}_2(\text{g}) \rightarrow 114\text{CO}_2(\text{g}) + 110\text{H}_2\text{O}(\text{l})$   
 a) 490 g                      b) 485 g                      c) 495 g                      d) 480 g
46. In an organic compound of molar mass  $108 \text{ g mol}^{-1}$  C, H and N atoms are present in 9: 1: 3.5 by weight. Molecular formula can be

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**Dixie Goss**



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## **Table of Contents Chemistry Section 1 Quiz Answers**

1. Understanding the eBook Chemistry Section 1 Quiz Answers
  - The Rise of Digital Reading Chemistry Section 1 Quiz Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Section 1 Quiz Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an eBook Platform
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Section 1 Quiz Answers
  - Personalized Recommendations
  - Chemistry Section 1 Quiz Answers User Reviews and Ratings
  - Chemistry Section 1 Quiz Answers and Bestseller Lists
5. Accessing Chemistry Section 1 Quiz Answers Free and Paid eBooks
  - Chemistry Section 1 Quiz Answers Public Domain eBooks
  - Chemistry Section 1 Quiz Answers eBook Subscription Services
  - Chemistry Section 1 Quiz Answers Budget-Friendly Options
6. Navigating Chemistry Section 1 Quiz Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Chemistry Section 1 Quiz Answers Compatibility with Devices
  - Chemistry Section 1 Quiz Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chemistry Section 1 Quiz Answers
  - Highlighting and Note-Taking Chemistry Section 1 Quiz Answers
  - Interactive Elements Chemistry Section 1 Quiz Answers
8. Staying Engaged with Chemistry Section 1 Quiz Answers

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Chemistry Section 1 Quiz Answers
- 9. Balancing eBooks and Physical Books Chemistry Section 1 Quiz Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chemistry Section 1 Quiz Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Chemistry Section 1 Quiz Answers
  - Setting Reading Goals Chemistry Section 1 Quiz Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemistry Section 1 Quiz Answers
  - Fact-Checking eBook Content of Chemistry Section 1 Quiz Answers
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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