

$$PV = nRT$$

0.0°C 1 atm  
22.414 L

## Gas Stoichiometry Practice Sheet

- 1) For the reaction  $2 \text{H}_{2(g)} + \text{O}_{2(g)} \rightarrow 2 \text{H}_2\text{O}_{(l)}$ , how many liters of water can be made from 5 L of oxygen gas and an excess of hydrogen?

$$\begin{array}{c|c|c|c} \cancel{5 \text{ L O}_2} & 1 \text{ mol O}_2 & 2 \text{ mol H}_2\text{O} & 22.414 \text{ L H}_2\text{O} \\ & 22.414 \text{ L O}_2 & 1 \text{ mol O}_2 & 1 \text{ mol H}_2\text{O} \end{array} = 10 \text{ L H}_2\text{O}$$

0.0°C  
1 atm

- 2) How many liters of water can be made from 55 grams of oxygen gas and an excess of hydrogen at STP?

$$\begin{array}{c|c|c|c} \cancel{55 \text{ g O}_2} & 1 \text{ mol O}_2 & 2 \text{ mol H}_2\text{O} & 22.414 \text{ L H}_2\text{O} \\ & 32 \text{ g O}_2 & 1 \text{ mol O}_2 & 1 \text{ mol H}_2\text{O} \end{array} = 77 \text{ L H}_2\text{O}$$

- 3) How many liters of water can be made from 55 grams of oxygen gas and an excess of hydrogen at a pressure of 12.4 atm and a temperature of 85°C?

$$\begin{array}{c|c|c|c} \cancel{55 \text{ g O}_2} & 1 \text{ mol O}_2 & 2 \text{ mol H}_2\text{O} & 22.414 \text{ L H}_2\text{O} \\ & 32 \text{ g O}_2 & 1 \text{ mol O}_2 & 1 \text{ mol H}_2\text{O} \end{array} = 3.4375 \text{ mol H}_2\text{O}$$

$$\text{O}_2 + 2\text{H}_2 \rightarrow 2\text{H}_2\text{O}$$

$$(12.4 \text{ atm}) V = (3.4375 \text{ mol H}_2\text{O}) \left( \frac{0.08205 \text{ L atm}}{\text{mol K}} \right) (358 \text{ K})$$

$$V = 8.14 \text{ L H}_2\text{O}$$

- 4) How many liters of water can be made from 34 grams of oxygen gas and an excess of hydrogen at STP?

$$\begin{array}{c|c|c|c} \cancel{34 \text{ g O}_2} & 1 \text{ mol O}_2 & 2 \text{ mol H}_2\text{O} & 22.414 \text{ L H}_2\text{O} \\ & 32 \text{ g O}_2 & 1 \text{ mol O}_2 & 1 \text{ mol H}_2\text{O} \end{array} = 23.8 \text{ L H}_2\text{O}$$

# Chemistry Stoichiometry Problem Sheet 2 Answer Key

**John A. Olmsted, Gregory M. Williams**



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