Market Street

Scale each of the following problems. Use correct units, and show your work for full credit

The specific heat of ethanol is 2.46 J/g°C. Find the heat energy required to rarse the temperature of 193 p of ethanol from 19°C to 39°C.

9 = (193)(2.46) (35-19) 19=75967

When a 120 g sample of atuminum absorbs 9612 J of heat energy, its temperature increases from 25°C to 115°C. Find the specific heat of aluminum, include the correct unit.

C = 7 = (120(15-25)

3. The specific heat of lead is 0.129 Jig*C. Find the amount of heat released when 2.4 mol of lead are cooled from SET SET OF HIS SECURITION.

9=(497/0.129)(22.5-3

How many J of energy are needed to raise the temperature of 165 mol of water from 10.85°C to 47.32°C?

16 Sule , 18.00 = 2970g

= (2970) (414) (47.32-10.55)

S. A 150 g sample of water (initially at 45.0°C) is mixed with a 200 g sample of water (initially at 64.0°C). First the final temperature of the system Coats.

(150)(4.18)(T=-45.0)=(200)(418)(T=-84.0) 6277- - 28215 = -8367_ +70224 /TE=6

A 25 g earnple of you (mitsely at 800.00°C) is dropped into 200 g of motor (mitsely at 30.00°C). The final temperature of the system is 40.22°C. Find the specific heat of iron.

- (25) (4.448) (40.22-800.00) = (200) (4.18) (40.22-30.00) 78999.5 C = 8543.9 Z C = 0.450 7/g °C

2. A east is sumple of mercury (specific feet = 0.14 Jrg °C, Indian ton Governor of 27.00°C) is placed into 134 g of water (initial temperature of the system.

 $(440)(0.14)(T_E - 22.00) = -(189)(4.18)(T_E - 35.00)$ $(1.6T_E - 1355.2 = -560.12T_E + 19604.2$

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Hongru Du

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