

# CHEMICAL QUANTITIES

# **Practice Problems**

On a piece of paper, solve the following problems.

YOU MUST SHOW YOUR WORK TO GET CREDIT FOR YOUR ANSWERS!!!

#### SECTION 10.1 THE MOLE: A MEASUREMENT OF MATTER

- What is the molar mass of sucrose (C<sub>12</sub>H<sub>22</sub>O<sub>11</sub>)?
- 2. What is the molar mass of each of the following compounds?
  - a, phosphorus pentachloride (PCI-)
  - b. uranium hexafluoride (UF<sub>4</sub>)
- 3. Calculate the molar mass of each of the following ionic compounds:
  - a. KMinO<sub>4</sub>
  - b. Cas(PO<sub>4</sub>):
- 4. How many moles is 3.52 × 1024 molecules of water?
- 5. How many atoms of zinc are in 0.60 mol of zinc?
- 6. What is the mass of 1.00 mol of exygen (O2)?

# SECTION 10.2 MOLE-MASS AND MOLE-VOLUME RELATIONSHIPS

- 1. What is the molar mass of each of the following compounds?
  - a. CaHagOa

b. NaHCOs.

e. Calling

d. KNH4SO<sub>4</sub>

- 2. Calculate the mass in grams of each of the following:
  - a. 8.0 mol lead oxide (PbO).

- d. 1.50 × 10<sup>-2</sup> mol molecular oxygen (O<sub>5</sub>)
- b. 0.75 mol hydrogen sulfide (H<sub>2</sub>S)

- e. 2.30 mol ethylene glycol (C<sub>2</sub>H<sub>2</sub>O<sub>2</sub>)
- e. 0.00100 mol silicon tetrahydride (SiH<sub>4</sub>).
- How many grams are in 1.73 mol of dinitrogen pentoxide (N<sub>2</sub>O<sub>2</sub>)?
- How many grams are in 0.658 mol of calcium phosphate [Ca<sub>2</sub>(PO<sub>4</sub>)<sub>2</sub>]?
- 5. Calculate the number of moles in each of the following:
  - a. 0.50 g sodium bromide (NaBr).

- d, 0.00100 g monochloromethane (CH<sub>2</sub>CI)
- b. 13.5 g magnesium nitrate [Mg(NO<sub>1</sub>)<sub>2</sub>]
- e. 1.50 × 10<sup>-3</sup> g propylene glycol [C<sub>3</sub>H<sub>4</sub>(OH)<sub>2</sub>]
- e. 1.02 g magnesium chloride (MgCl<sub>2</sub>)
- 6. A chemist plans to use 435.0 grams of ammonium nitrate (NH<sub>4</sub>NO<sub>5</sub>) in a reaction. How many moles of the compound is this?
- 7. A solution is to be prepared in a laboratory. The solution requires 0.0465 mol of quinine (C<sub>20</sub>H<sub>24</sub>N<sub>2</sub>O<sub>2</sub>). What mass, in grams, should the laboratory technician obtain in order to make the solution?
- 8. What is the volume at STP of 2.66 mol of methane (CHa) gas?
- 9. How many moles is 135 L of ammonia (NH<sub>3</sub>) gas at STP?

# **Chemistry Chapter 10 Study Guide Answers**

Susan McMurry, John McMurry

## **Chemistry Chapter 10 Study Guide Answers:**

Study Guide and Solutions Manual for McMurry's Organic Chemistry, Fifth Edition Susan McMurry,2000 Provides answers and explanations to all in text and end of chapter exercises Also includes summaries of name reactions functional group synthesis and reactions lists of reagents and abbreviations and articles on topics ranging from infrared absorption frequencies to the Nobel Price winners in Chemistry This edition now includes all new artwork expanded in text problems summary quizzes approximately every three chapters more detailed explanations in solutions and chapter outlines

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