

## Stoichiometry Worksheet # 1

- 1) Glucose is used as a source of energy by the human body. The overall reaction in the body is



Calculate the number of moles of oxygen needed to oxidize 12.5 g of  $\text{C}_6\text{H}_{12}\text{O}_6$ .

- 2) Ammonia is synthesized from hydrogen and nitrogen according to the following equation.



If an excess of nitrogen is reacted with 3.41 moles of hydrogen gas, how many grams of ammonia ( $\text{NH}_3$ ) can be produced?

- 3) Assume that in the decomposition of potassium chlorate,  $\text{KClO}_3$ ,



80.5 moles of  $\text{O}_2$  form. How many grams of potassium chlorite, the other product would be formed?

- 4) In a single displacement reaction,



9.23 moles of aluminum react with excess hydrochloric acid. How many moles of hydrogen will be produced?

- 5) The compound "Cisplatin"  $\text{PtCl}_2(\text{NH}_3)_2$ , has been found to be effective in treating some types of cancer. It can be synthesized using the following reaction



A. How many moles of "cisplatin" can be produced from 2.50 g  $\text{K}_2\text{PtCl}_4$ ?

B. How many moles of  $\text{NH}_3$  would be needed if 3.8 moles of  $\text{K}_2\text{PtCl}_4$  were present?

# Chemistry Chapter 12 Stoichiometry Test

**Thomas Evangelist**



## **Chemistry Chapter 12 Stoichiometry Test:**

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**Essential Concepts of Chemistry** Sharon J. Sherman, 1998-09

**Teaching Science for Understanding** James Joseph Gallagher, 2007 Offers middle and high school science teachers practical advice on how they can teach their students key concepts while building their understanding of the subject through various levels of learning activities

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## **Chemistry Chapter 12 Stoichiometry Test** Book Review: Unveiling the Magic of Language

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