

12

STOICHIOMETRY

Practice Problems

In your notebook, solve the following problems.

SECTION 12.1 THE ARITHMETIC OF EQUATIONS

Use the 3-step problem-solving approach you learned in Chapter 1.

1. An apple pie needs 10 large apples, 2 cups (top and bottom), and 1 tablespoon of cinnamon. Write a balanced equation that fits this situation. How many apples are needed to make 20 pies?
2. Two moles of potassium chloride and three moles of oxygen are produced from the decomposition of two moles of potassium chlorate, $2\text{KClO}_3(\text{s})$. Write the balanced equation. How many moles of oxygen are produced from 12 moles of potassium chlorate?
3. Using the equation from problem 2, how many moles of oxygen are produced from 14 moles of potassium chlorate?
4. Two molecules of hydrogen react with one molecule of oxygen to produce two molecules of water. How many molecules of water are produced from 3.8×10^{23} molecules of oxygen? How many moles of water are produced from 22.5 moles of oxygen?

SECTION 12.2 CHEMICAL CALCULATIONS

1. Calculate the number of moles of hydrogen chloride produced from 10 moles of hydrogen.



2. Calculate the number of moles of chlorine needed to form 14 moles of iron(II) chloride.



3. Calculate the number of grams of nitrogen dioxide that are produced from 4 moles of nitric oxide.



4. Calculate the mass of oxygen produced from the decomposition of 75.0 g of potassium chlorate.



5. Calculate the mass of silver needed to react with chlorine to produce 94 g of silver chloride. Write a balanced equation first.

6. How many liters of carbon monoxide at STP are needed to react with 4.00 g of oxygen gas to produce carbon dioxide?



7. Calculate the number of liters of oxygen gas needed to produce 15.0 liters of dinitrogen monoxide. Assume all gases are at the same conditions of temperature and pressure.



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Jiyuan Zhang



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1. Understanding the eBook Chemistry Chapter 12 Stoichiometry Worksheet Answer Key
 - The Rise of Digital Reading Chemistry Chapter 12 Stoichiometry Worksheet Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Chapter 12 Stoichiometry Worksheet Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Chapter 12 Stoichiometry Worksheet Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Chapter 12 Stoichiometry Worksheet Answer Key
 - Personalized Recommendations
 - Chemistry Chapter 12 Stoichiometry Worksheet Answer Key User Reviews and Ratings
 - Chemistry Chapter 12 Stoichiometry Worksheet Answer Key and Bestseller Lists
5. Accessing Chemistry Chapter 12 Stoichiometry Worksheet Answer Key Free and Paid eBooks
 - Chemistry Chapter 12 Stoichiometry Worksheet Answer Key Public Domain eBooks
 - Chemistry Chapter 12 Stoichiometry Worksheet Answer Key eBook Subscription Services
 - Chemistry Chapter 12 Stoichiometry Worksheet Answer Key Budget-Friendly Options

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 - ePub, PDF, MOBI, and More
 - Chemistry Chapter 12 Stoichiometry Worksheet Answer Key Compatibility with Devices
 - Chemistry Chapter 12 Stoichiometry Worksheet Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Chapter 12 Stoichiometry Worksheet Answer Key
 - Highlighting and Note-Taking Chemistry Chapter 12 Stoichiometry Worksheet Answer Key
 - Interactive Elements Chemistry Chapter 12 Stoichiometry Worksheet Answer Key
8. Staying Engaged with Chemistry Chapter 12 Stoichiometry Worksheet Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Chapter 12 Stoichiometry Worksheet Answer Key
9. Balancing eBooks and Physical Books Chemistry Chapter 12 Stoichiometry Worksheet Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Chapter 12 Stoichiometry Worksheet Answer Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chemistry Chapter 12 Stoichiometry Worksheet Answer Key
 - Setting Reading Goals Chemistry Chapter 12 Stoichiometry Worksheet Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemistry Chapter 12 Stoichiometry Worksheet Answer Key
 - Fact-Checking eBook Content of Chemistry Chapter 12 Stoichiometry Worksheet Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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