Balancing Equations and Simple Stoichiometry-KEY

Balance the following equations:

- 1 N₂ + 3 F₂ → 2 NF₃
- 2) 2 C₆H₁₀ + 17 O₂ → 12 CO₂ + 10 H₂O
- 3) 1 HBr + 1 KHCO₃ → 1 H₂O + 1 KBr + 1 CO₂
- 4) 2 GaBr₃ + 3 Na₂SO₃ → 1 Ga₂(SO₃)₃ + 6 NaBr
- 5) 3 SnO + 2 NF₃ → 3 SnF₂ + 1 N₂O₃

Using the following equation:

- 6) How many grams of sodium sulfate will be formed if you start with 200 grams of sodium hydroxide and you have an excess of sulfuric acid?
- ? g Na₂SO₄ = 200 g NaOH→mol NaOH→mol Na₂SO₄ → g Na₂SO₄

Using the following equation:

$$Pb(SO_4)_2 + 4 LiNO_3 \rightarrow Pb(NO_3)_4 + 2 Li_2SO_4$$

How many grams of lithium nitrate will be needed to make 250 grams of lithium sulfate, assuming that you have an adequate amount of lead (IV) sulfate to do the reaction?

Molar masses: LiNO₃ = 68.95 g/mol & Li₂SO₄ = 109.9 g/mol ? g LiNO₃ = 250 g Li₂SO₄ \rightarrow mol Li₂SO₄ \rightarrow mol LiNO₃ \rightarrow g LiNO₃

$$250 \text{ g Li}_2SO_4 \times \underline{\text{I mol Li}_2SO_4} = 2.27 \text{ mol Li}_2SO_4 \times \underline{4 \text{ mol LiNO}_3} = 4.54 \text{ mol LiNO}_3$$

 $109.9 \text{ g Li}_2SO_4 = 2.27 \text{ mol Li}_2SO_4 \times \underline{4 \text{ mol LiNO}_3} = 4.54 \text{ mol LiNO}_3$

<u>Chemistry Chapter 1stoichiometry Worksheet Answer</u> <u>Key</u>

B. Richard Siebring, Mary Ellen Schaff

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