

Answers to Chapter 3 Study Questions

- $3(12.0) + 8(1.01) + 3(16.0) = 92.1 \text{ g/mole}$ b) 92.1 g c) $6.02 \times 10^{23} \text{ molecules}$
 - $0.217 \text{ moles} \times \frac{92.1 \text{ g}}{1 \text{ mole}} = 20.0 \text{ g}$
- 4 atoms (one N + 3 H)
 - $1 \text{ mole } \text{NH}_3 \times \frac{6.02 \times 10^{23} \text{ molecules}}{1 \text{ mole } \text{NH}_3} \times \frac{4 \text{ atoms}}{1 \text{ molecule}} = 2.41 \times 10^{24} \text{ atoms}$
 - $3.40 \text{ g } \text{NH}_3 \times \frac{1 \text{ mole } \text{NH}_3}{17.0 \text{ g } \text{NH}_3} \times \frac{6.02 \times 10^{23} \text{ molecules}}{1 \text{ mole } \text{NH}_3} \times \frac{4 \text{ atoms}}{1 \text{ molecule}} = 4.82 \times 10^{23} \text{ atoms}$
- Molar mass of $\text{NaNO}_2 = 23.0 + 14.0 + 2(16.0) = 69.0 \text{ g/mole}$
 $\% \text{ Na} = 23.0/69.0 = 33.3\% \text{ Na}$; $\% \text{ N} = 14.0/69.0 = 20.3\% \text{ N}$; $\% \text{ O} = 2(16.0)/69.0 = 46.4\% \text{ O}$
33.3% Na, 20.3% N and 46.4% O.
- In 100 g of this compound, there are 40.7 g C, 5.1 g H, and 54.2 g O
 $40.7 \text{ g C} \times \frac{1 \text{ mole C}}{12.0 \text{ g C}} = 3.39 \text{ moles C}$ $3.39/3.39 = 1 \times 2 = 2$
 $5.1 \text{ g H} \times \frac{1 \text{ mole H}}{1.0 \text{ g H}} = 5.1 \text{ moles H}$ $5.1/3.39 = 1.5 \times 2 = 3$
 $54.2 \text{ g O} \times \frac{1 \text{ mole O}}{16.0 \text{ g O}} = 3.39 \text{ moles O}$ $3.39/3.39 = 1 \times 2 = 2$
empirical formula = $\text{C}_2\text{H}_3\text{O}_2$
 - Molar mass of $\text{C}_2\text{H}_3\text{O}_2 = 2(12.0) + 3(1.0) + 2(16.0) = 59.0 \text{ g/mole}$
 $118/59.0 = 2 \rightarrow$ molecular formula = **$\text{C}_4\text{H}_6\text{O}_4$**
- In 25.0 g of compound, there are 7.20 g Mg, 3.55 g C and $25.0 - (7.20 + 3.55) = 14.25 \text{ g O}$.
 $7.20 \text{ g Mg} \times \frac{1 \text{ mole Mg}}{24.3 \text{ g Mg}} = 0.296 \text{ moles Mg}$ $0.296/0.296 = 1$
 $3.55 \text{ g C} \times \frac{1 \text{ mole C}}{12.0 \text{ g C}} = 0.296 \text{ moles C}$ $0.296/0.296 = 1$
 $14.25 \text{ g O} \times \frac{1 \text{ mole O}}{16.0 \text{ g O}} = 0.891 \text{ moles O}$ $0.891/0.296 = 3$
formula = MgCO_3
 - $\% \text{ Mg} = 7.20/25.0 = 28.8\% \text{ Mg}$; $\% \text{ C} = 3.55/25.0 = 14.2\% \text{ C}$; $\% \text{ O} = 14.25/25.0 = 57.0\% \text{ O}$. (You should get the same result using molar mass and atomic masses.)
 - $13.9 \text{ g compound} \times \frac{28.8 \text{ g Mg}}{100 \text{ g compound}} = 4.00 \text{ g Mg}$

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Developing Language and Literacy in English across the Secondary School Curriculum Urszula Clark, 2018-07-21 This book draws on original research and a language based pedagogy approach to examine how secondary schools in the UK can devise and implement coherent language and literacy across curriculum policies and strategies so that grammar and associated metalanguage becomes an integral part of their day to day curriculum practices The research was undertaken in three 11 to 18 secondary schools in England where the majority of students are categorised as having English as a second language EAL and where a significant minority are also socially disadvantaged in two of the three The author argues that paying explicit attention to the linguistic structures through which subject knowledge is realised can be of benefit to all pupils in ways that are also socially just and democratic This book provides an important bridge between academic theory and educational practice that will appeal to applied linguists and sociolinguists as well as to teachers teacher trainers and practitioners

Journal of Biological Education , 1989

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