

Section 6.1: Organizing the Elements

- Group 1 on the periodic table is called **Alkali Metals**
- Group 2 on the periodic table is called **Alkaline Earth Metals**
- Elements in groups 3-12 are called **Transition Metals**
- Group 17 on the periodic table is called **Halogens**
- Group 18 on the periodic table is called **Noble Gases**
- Each period on the periodic table corresponds to a **principal energy level**
- The 2 series that make up the inner transition metals are **Lanthanides** and **Actinides**
- Complete the following chart:

Element	Symbol	Group #	Period #	Group Name	Metal, Metalloid, Nonmetal
Zirconium	Zr	4	5	Transition Metals	Metal
Radium	Ra	2	7	Alkaline Earth Metals	Metal
Bismuth	Bi	15	6	Nitrogen Group	Metal
Gallium	Ga	13	4	Boron Group	Metal

- What are triads? **Sets of 3 elements with similar properties** Whose idea were they? **Dobereiner**
- The modern periodic table is arranged by increasing **atomic number**. Whose idea was this? **Moseley**
- Mendeleev** developed the first periodic table, which was arranged by increasing **atomic mass**
- Determine which of the following pairs will have similar properties (Yes or no):
 - Ca and Be? **Yes – same group (Group 2 – Alkaline Earth Metals)**
 - Sr and Y? **No – not the same group**
 - C and Ge? **Yes – same group (Group 14 – Carbon Group)**

Section 6.3: Periodic Trends

- As you go across a period, the atomic radius (size) **decreases** because **the attraction between protons and electrons increases, which makes the atom smaller (opposite of tug of war)**
- As you go down a group, the atomic radius (size) **increases** because **the number of principal energy levels increases**
- Circle the atom with the largest atomic radius and put a square around the atom with the smallest radius:

a.	<div style="border: 1px solid black; border-radius: 50%; padding: 2px;">Li</div>	<div style="border: 1px solid black; padding: 2px;">Ne</div>	F	B
b.	<div style="border: 1px solid black; padding: 2px;">N</div>	<div style="border: 1px solid black; border-radius: 50%; padding: 2px;">Bi</div>	S	P
c.	Na	<div style="border: 1px solid black; border-radius: 50%; padding: 2px;">Sr</div>	Al	<div style="border: 1px solid black; padding: 2px;">Cl</div>
- Ionization energy is the energy required to **remove** an electron.

Chemistry Chapter 6 1 Organizing The Elements

Worksheet Answers

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Table of Contents Chemistry Chapter 6 1 Organizing The Elements Worksheet Answers

1. Understanding the eBook Chemistry Chapter 6 1 Organizing The Elements Worksheet Answers
 - The Rise of Digital Reading Chemistry Chapter 6 1 Organizing The Elements Worksheet Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Chapter 6 1 Organizing The Elements Worksheet Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an eBook Chemistry Chapter 6 1 Organizing The Elements Worksheet Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Chapter 6 1 Organizing The Elements Worksheet Answers
 - Personalized Recommendations

- Chemistry Chapter 6 1 Organizing The Elements Worksheet Answers User Reviews and Ratings
- Chemistry Chapter 6 1 Organizing The Elements Worksheet Answers and Bestseller Lists
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- 6. Navigating Chemistry Chapter 6 1 Organizing The Elements Worksheet Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemistry Chapter 6 1 Organizing The Elements Worksheet Answers Compatibility with Devices
 - Chemistry Chapter 6 1 Organizing The Elements Worksheet Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Chapter 6 1 Organizing The Elements Worksheet Answers
 - Highlighting and Note-Taking Chemistry Chapter 6 1 Organizing The Elements Worksheet Answers
 - Interactive Elements Chemistry Chapter 6 1 Organizing The Elements Worksheet Answers
- 8. Staying Engaged with Chemistry Chapter 6 1 Organizing The Elements Worksheet Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Chapter 6 1 Organizing The Elements Worksheet Answers
- 9. Balancing eBooks and Physical Books Chemistry Chapter 6 1 Organizing The Elements Worksheet Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Chapter 6 1 Organizing The Elements Worksheet Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chemistry Chapter 6 1 Organizing The Elements Worksheet Answers
 - Setting Reading Goals Chemistry Chapter 6 1 Organizing The Elements Worksheet Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemistry Chapter 6 1 Organizing The Elements Worksheet Answers
 - Fact-Checking eBook Content of Chemistry Chapter 6 1 Organizing The Elements Worksheet Answers

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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