

7

CHEMICAL QUANTITIES

Quiz for CHAPTER 7

Fill in the word(s) that will make each statement true.

1. Counting, finding mass, and finding volume are all ways of measuring the amount of 1. 1. _____ 7.1
2. The 2 is the mass, in grams, of one mole of atoms of a monatomic element. 2. _____ 7.1
3. The term 3 may be used in place of the term gram formula mass to refer to the mass of any element or compound. 3. _____ 7.2
4. The representative particle of most elements is the 4. 4. _____ 7.2
5. To convert moles of oxygen gas (O_2) to mass in grams, you can use the conversion factor 5. 5. _____ 7.2

Classify each of these statements as always true, AT; sometimes true, ST; or never true, NT.

- _____ 6. The density of the gas C_2H_6 at STP is 1.34 g/L. 7.2
- _____ 7. The volume of 3 moles of oxygen at STP is 67.2 L. 7.2
- _____ 8. One mole of a gas occupies 22.4 L. 7.2
- _____ 9. To convert a volume of 3.20 L of hydrogen gas to the number of moles, you would use the conversion factor $\frac{22.4 \text{ L}}{1.00 \text{ mol } H_2}$. 7.2
- _____ 10. The density of a gas at STP can be used to calculate the gram formula mass of that gas. 7.2

Solve these problems and show your work in the space provided.

11. What mass of gold contains as many atoms as 3.05 g of silver? 7.2

12. What is the empirical formula of a compound that is 40.7% carbon, 54.2% oxygen, and 5.1% hydrogen? 7.3

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