

PHYSICAL EQUILIBRIUM

- 1. The chemical reaction which takes place in a closed system is called equilibrium reaction.
- 2. It is a reversible reaction that both forward and backward reactions.
- 3. Equilibrium is a state where rate of forward reaction is equal to rate of backward reaction.
- 4. Equilibrium is achieved only in a closed system.
- 5. **Equilibrium** is a state of forward and backward reaction.
- 6. In equilibrium, the composition of reactants and products remains constant.

GRAPHICAL REPRESENTATIONS



1

For the equilibrium, $A \rightleftharpoons B$, the reaction of the rate of the forward (a) & backward (b) reaction with time is given by



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2

Which of the given statements does not describe the equilibrium state properly?

- (A) The equilibrium can be approached from either direction.
- (B) The equilibrium can be imposed only if the system is an isolated system.
- (C) The free energy change of constant pressure and temperature is zero.
- (D) It is dynamic in nature.

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3

In the given reaction, $A + 2B \rightleftharpoons C$, 2 moles each of A & B are present in 10 L of reaction container at time, T kept at 10°C . Calculate K for the reaction.

- (A) 1/3 (B) 1/6 (C) 1/12 (D) 1/2

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4

In a given temperature, the equilibrium constant for the reaction,



and H_2 and O_2 respectively. If K_1 is 4×10^2 , then K_2 will be

- (A) 4×10^2 (B) $1/4 \times 10^2$ (C) 4×10^4 (D) $1/4 \times 10^4$

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For the reaction,



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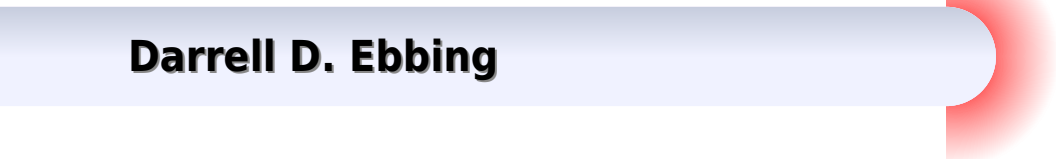
Which one of the following conditions will favour maximum formation of the product in the reaction $\text{A} + \text{B} \rightleftharpoons \text{C} + \text{D}$ $\Delta H = -10\text{ kJ}$

- (A) Low temperature and high pressure (B) High temperature and high pressure
- (C) Low temperature and low pressure (D) High temperature and low pressure

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Darrell D. Ebbing



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