

**CHAPTER 6 REVIEW***Chemical Bonding***MIXED REVIEW**

**SHORT ANSWER** Answer the following questions in the space provided.

1. Name the type of energy that is a measure of strength for each of the following types of bonds:

|   |                         |
|---|-------------------------|
| _____ <b>lattice energy</b> _____           | <b>a.</b> ionic bond    |
| _____ <b>bond energy</b> _____              | <b>b.</b> covalent bond |
| _____ <b>enthalpy of vaporization</b> _____ | <b>c.</b> metallic bond |

2. Use the electronegativity values shown in **Figure 20**, on page 161 of the text, to determine whether each of the following bonds is nonpolar covalent, polar covalent, or ionic.

|                                   |                 |                                      |               |
|-----------------------------------|-----------------|--------------------------------------|---------------|
| _____ <b>ionic</b> _____          | <b>a.</b> H—F   | _____ <b>nonpolar covalent</b> _____ | <b>d.</b> H—H |
| _____ <b>ionic</b> _____          | <b>b.</b> Na—Cl | _____ <b>polar covalent</b> _____    | <b>e.</b> H—C |
| _____ <b>polar covalent</b> _____ | <b>c.</b> H—O   | _____ <b>polar covalent</b> _____    | <b>f.</b> H—N |

3. How is a hydrogen bond different from an ionic or covalent bond?

A hydrogen bond is a dipole-dipole attraction between a partially positive hydrogen  
atom and the unshared electron pair of a strongly electronegative atom such as O,  
N, or F. Unlike ionic or covalent bonds, in which electrons are given up or shared,  
the hydrogen bond is a weaker attraction. Hydrogen bonds are generally  
intermolecular, while ionic and covalent bonds occur between ions or atoms  
respectively.

4. H<sub>2</sub>S and H<sub>2</sub>O have similar structures and their central atoms belong to the same group. Yet H<sub>2</sub>S is a gas at room temperature and H<sub>2</sub>O is a liquid. Use bonding principles to explain why this is.

Oxygen has higher electronegativity than sulfur, which creates a highly polar bond.  
Increased polarity in H<sub>2</sub>O bonds means a stronger intermolecular attraction, making  
water a liquid at room temperature. Hydrogen bonding exists between water  
molecules, but not between hydrogen sulfide molecules.

# Chemistry Covalent Bonding Study Guide Answers

**David R. Klein, Laurie S. Starkey**



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## **Table of Contents Chemistry Covalent Bonding Study Guide Answers**

1. Understanding the eBook Chemistry Covalent Bonding Study Guide Answers
  - The Rise of Digital Reading Chemistry Covalent Bonding Study Guide Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Covalent Bonding Study Guide Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chemistry Covalent Bonding Study Guide Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Covalent Bonding Study Guide Answers
  - Personalized Recommendations
  - Chemistry Covalent Bonding Study Guide Answers User Reviews and Ratings

- Chemistry Covalent Bonding Study Guide Answers and Bestseller Lists
- 5. Accessing Chemistry Covalent Bonding Study Guide Answers Free and Paid eBooks
  - Chemistry Covalent Bonding Study Guide Answers Public Domain eBooks
  - Chemistry Covalent Bonding Study Guide Answers eBook Subscription Services
  - Chemistry Covalent Bonding Study Guide Answers Budget-Friendly Options
- 6. Navigating Chemistry Covalent Bonding Study Guide Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Chemistry Covalent Bonding Study Guide Answers Compatibility with Devices
  - Chemistry Covalent Bonding Study Guide Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chemistry Covalent Bonding Study Guide Answers
  - Highlighting and Note-Taking Chemistry Covalent Bonding Study Guide Answers
  - Interactive Elements Chemistry Covalent Bonding Study Guide Answers
- 8. Staying Engaged with Chemistry Covalent Bonding Study Guide Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chemistry Covalent Bonding Study Guide Answers
- 9. Balancing eBooks and Physical Books Chemistry Covalent Bonding Study Guide Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chemistry Covalent Bonding Study Guide Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Chemistry Covalent Bonding Study Guide Answers
  - Setting Reading Goals Chemistry Covalent Bonding Study Guide Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemistry Covalent Bonding Study Guide Answers
  - Fact-Checking eBook Content of Chemistry Covalent Bonding Study Guide Answers
  - Distinguishing Credible Sources

13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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planning select a theme choose a sunday

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