
Bond Energy Study Guide (Chemistry 2.2)

1. Why atoms form bonds

Atoms form bonds to reach a lower, more stable energy state.

- Each atom has **protons (+)** in its nucleus and **electrons (-)** around it.
- **Coulombic attraction** is the force between opposite charges. When atoms get close, their nuclei attract the other atom's electrons.
- This attraction lowers potential energy — that's why bonds form.

When the bond forms:

- **Energy is released** → the system becomes more stable.
- **Bond making = exothermic process.**

When a bond breaks:

- **Energy is absorbed** → you must overcome that attraction.
- **Bond breaking = endothermic process.**

2. Endothermic vs. Exothermic

Type of Process	Energy Flow	What Happens	ΔH (Enthalpy Change) Sign
Endothermic	Energy absorbed	Bonds are broken	Positive (+)
Exothermic	Energy released	Bonds are formed	Negative (-)

Example:

- When hydrogen and chlorine form HCl, the reaction releases energy — **exothermic**.
- When you separate HCl into H and Cl atoms, you need energy — **endothermic**.

3. Net Enthalpy Change (ΔH)

Chemistry Energy Study Guide Key

A Gutmann



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