

12U Equilibrium Unit Test

Full Name: _____

Date: _____

Multiple Choice K/U

Identify the letter of the choice that best completes the statement or answers the question.

- ____ 1. When solid lead(II) phosphate is in equilibrium with its ions, the ratio of lead(II) ions to phosphate ions is which of the following?
- 1:1
 - 1:2
 - 2:1
 - 2:3
 - 3:2
- ____ 2. In a saturated solution of silver phosphate, the concentration of silver ion is 4.5×10^{-4} mol/L. The K_{sp} of silver phosphate would be which of the following?
- 6.8×10^{-8}
 - 1.4×10^{-14}
 - 1.5×10^{-13}
 - 1.0×10^{-13}
 - none of the above
- ____ 3. For the equilibrium system below, which of the following would result in an increase in the quantity of $\text{PCl}_5(\text{g})$?
- $$\text{PCl}_5(\text{g}) + \text{Cl}_2(\text{g}) \rightleftharpoons \text{PCl}_3(\text{g}) + 45 \text{ kJ}$$
- increasing temperature
 - increasing the size of the container
 - decreasing temperature
 - removing some $\text{Cl}_2(\text{g})$
 - injecting some He gas
- ____ 4. 10.0 mol of ammonia gas is injected into a 4.0 L container. At equilibrium 1.2 mol of hydrogen gas is found in the container. The number of moles of ammonia gas left in the container must be which of the following?
- $$\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \rightleftharpoons 2\text{NH}_3(\text{g}) + 94 \text{ kJ}$$
- 6.8
 - 6.4
 - 8.6
 - 8.8
 - 9.2
- ____ 5. A closed flask contains this equilibrium
- $$2\text{NO}_2(\text{g}) \text{ (brown gas)} \rightleftharpoons \text{N}_2\text{O}_4(\text{g}) \text{ (colourless gas)}$$
- When the equilibrium is heated at a constant pressure it gets browner. This must mean which of the following?
- K_c is very large
 - it is endothermic
 - it is exothermic
 - both a and b
 - both a and c
- ____ 6. Which equilibrium shows a decrease in [products] when the volume increases?
- $2\text{H}_2(\text{g}) + \text{O}_2(\text{g}) \rightleftharpoons 2\text{H}_2\text{O}(\text{g})$
 - $\text{Cl}_2(\text{g}) + \text{PCl}_5(\text{g}) \rightleftharpoons \text{PCl}_3(\text{g})$
 - $\text{H}_2(\text{g}) + \text{I}_2(\text{g}) \rightleftharpoons 2\text{HI}(\text{g})$
 - $2\text{NH}_3(\text{g}) \rightleftharpoons \text{N}_2 + 3\text{H}_2(\text{g})$
- I
 - I and II
 - IV
 - III and IV
 - II
- ____ 7. Consider this equilibrium
- $$4\text{HCl}(\text{g}) + \text{O}_2(\text{g}) \rightleftharpoons 2\text{H}_2\text{O}(\text{g}) + 2\text{Cl}_2(\text{g})$$
- The equilibrium law expression for the balanced chemical equation would be
- $[\text{HCl}][\text{O}_2] / [\text{H}_2\text{O}][\text{Cl}_2]$
 - $[\text{H}_2\text{O}]^2[\text{Cl}_2]^2 / [\text{HCl}]^4[\text{O}_2]$
 - $2[\text{H}_2\text{O}][\text{Cl}_2] / 4[\text{HCl}][\text{O}_2]$
 - $[\text{HCl}]^4[\text{O}_2] / [\text{H}_2\text{O}]^2[\text{Cl}_2]^2$
 - $[\text{H}_2\text{O}][\text{Cl}_2] / \text{HCl}[\text{O}_2]$
- ____ 8. The K_{sp} for a substance A_2B is 1.2×10^{-12} . The $[\text{B}]$ must be which of the following?
- 4.0×10^{-11}
 - 3.6×10^{-12}
 - 6.7×10^{-5}
 - 1.3×10^{-4}
 - 2.6×10^{-4}

Chemistry Equilibrium Multiple Choice Practice Test

Jo Ann Robinson, Jacqueline E. Taylor



Chemistry Equilibrium Multiple Choice Practice Test:

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