

CHEMISTRY: REVIEW FOR 1ST SEMESTER FINAL (#1)

CH 2 Measurement and Calculations

$$K = ^\circ C + 273$$

$$\text{Density} = \text{Mass/Volume}$$

$$^\circ F = 1.8^\circ C + 32$$

- 1a) Temperature is $-40^\circ C$. What is it in Kelvin?
b) What is $25^\circ C$ on the Fahrenheit scale?
c) What is $-32^\circ C$ on the Fahrenheit scale?
d) Temperature is $150^\circ F$. What is it in Celsius?
e) Temperature is $548 K$. What is it in Celsius?
f) Temperature is $-125^\circ C$. What is it in Kelvin?
2. Ethanol has a density of 0.785 g/mL . Find the volume of 120.2 g ethanol.
3. A 224.4 g sample of mercury occupies a volume of 16.5 mL . What is the density of mercury?
4. A block measures 6 cm by 3 cm by 4 cm .
 - a) Find the volume of the block.
 - b) Calculate the mass if the density of the block is 2.82 g/cm^3
- 5a) $1 \text{ cm}^3 = \underline{\hspace{2cm}} \text{ mL}$
b) The unit for volume of solids is
c) The common unit for volume of liquids is .
6. A 340 g metal displaces 155 mL when dropped into a large graduated cylinder. What is its density?
7. The lowest possible temperature is $-273^\circ C$.
 - a) What is the temperature in Kelvin?
 - b) This temperature is known as the
8. A 52.4 g of iron is added to 45.0 mL of water in a graduated cylinder. The water level rises to 51.7 mL . What is the density of iron?
9. A 5.25 g lead block is dropped into a graduated cylinder containing 10.4 mL of water. The water level rises to 10.9 mL . What is the density of the block?

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