

(vi) In the reaction $I_2 + I^- \longrightarrow I_3^-$, which is the Lewis base?

(a) I_2

(b) I^-

(c) I_3^-

(d) None of these

(vii) In what manner will increase of pressure affect the following equilibrium?



(a) No effect

(b) Shift in the forward direction

(c) Shift in the reverse direction

(d) Increase in the field of H_2

(viii) A Metal ion M^{3+} loses 3 electrons, its oxidation number will be

(a) +3

(b) -3

(c) 0

(d) +6

(ix) The oxidation states of iodine in HIO_4 , H_3IO_5 and H_5IO_6 are respectively

(a) +1, +3, +7

(b) +7, +7, +3

(c) +7, +7, +7

(d) +7, +5, +3

(x) $:CCl_2$ is

(a) an electrophile

(b) a free radical

(c) a nucleophile

(d) none of these.

SECTION B

Q 2. Answer the following questions.

(i) Write the general electronic configuration of p and d block elements.

(ii) At $0^\circ C$, ice and water are present in equilibrium. What will happen on increasing the pressure.

(iii) Define oxidation and reduction according to classical concept.

(iv) What are Carbocations and Free radicals?

(v) Give two methods of preparation of Alkanes?

(vi) Why is benzene extraordinarily stable though it contains three double bonds?

(vii) Define Inductive effect & Hyper-conjugation.

(viii) Define

a) Resonance

b) Ozonolysis

(ix) Define Sulphonation. Give Example.

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