

Chapter 7

Quantities in Chemical Reactions

Solutions for Practice Problems

Student Textbook page 237

1. Problem

Consider the following reaction.



- (a) Write the ratio of H_2 molecules: O_2 molecules: H_2O molecules.
- (b) How many molecules of O_2 are required to react with 100 molecules of H_2 , according to your ratio in part (a)?
- (c) How many molecules of water are formed when 2478 molecules of O_2 react with H_2 ?
- (d) How many molecules of H_2 are required to react completely with 6.02×10^{23} molecules of O_2 ?

What Is Required?

- (a) You need to give the ratio of molecules in the reaction.
- (b) You need to find the number of O_2 molecules that will react with the given number of H_2 molecules.
- (c) You need to find the number of water molecules that will form from the given number of reactants.
- (d) You need to find the number of H_2 molecules that will react with the given number of O_2 molecules.

What Is Given?

The balanced equation for the reaction is given. The number of molecules of reactants or products is given in each question.

Plan Your Strategy

- (a) The ratio is given by the whole number in front of each reactant and product in the balanced equation.
- (b) Equate the ratio of number of molecules to the whole number ratio obtained in (a) for hydrogen:oxygen, and solve for the number of oxygen molecules.
- (c) As in (b), equate the ratio of the number of molecules to the whole number ratio obtained in (a) for oxygen:water. Solve for the number of water molecules.
- (d) As in (b) and (c), equate the ratio of the number of molecules to the whole number ratio obtained in (a) for hydrogen:oxygen. Solve for the number of hydrogen molecules.

Chemistry Packet Chapter Chemical Quantities Practice Problems Answers

**Donald J. Wink, Sharon Fetzer-
Gislason, Sheila McNicholas**



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