

1. Determine the kinds of intermolecular forces (dispersion, dipole-dipole, H-bonding) that are present in each of the following elements or compounds.

- a. Kr dispersion
- b. NCl_3 dispersion, dipole-dipole
- c. SiH_4 dispersion
- d. HF dispersion, dipole-dipole, H-bonding
- e. N_2 dispersion
- f. NH_3 dispersion, dipole-dipole, H-bonding
- g. CO dispersion, dipole-dipole
- h. CO_2 dispersion,

2. Arrange the following compounds in order of increasing boiling point, explain your reasoning.



3. Which one of the molecules would require the most energy to break the IMFs? How would you rank them in term of their boiling points?






4. Which one of the following compounds is more likely to be a liquid at room temperature, and why?



While all three compounds have similar molar mass, and all have dispersion and dipole-dipole forces, only hydrogen peroxide has H-bonding, thus likely to have the highest b.p. amongst the three compounds.

5. The surface tension and viscosity values for diethyl ether, acetone, ethanol and ethylene glycol are shown here:

Compound	Molecule	Surface Tension (mN/m)	Viscosity (mPa s)
diethyl ether $\text{C}_2\text{H}_5\text{OC}_2\text{H}_5$		17	0.22
ethanol $\text{C}_2\text{H}_5\text{OH}$		22	1.07
ethylene glycol $\text{CH}_2(\text{OH})\text{CH}_2(\text{OH})$		48	16.1

- a. Explain the differences in viscosity in terms of the size and shape of the molecules and their IMFs.
The molecules with the most IMFs will have the highest viscosity. Ethylene glycol is the largest molecule and has two H-bonding sites, whereas ethanol is small and has one H-bonding. Diethylether is the least viscous since it does not have H-bonding.
- b. Explain their differences in surface tension in terms of the size and shape of the molecules and their IMFs.
Similar trend is observed for surface tension, where the molecule with the most IMFs (ethylene glycol) will exhibit the highest surface tension.

Chemistry Practice Test Intermolecular Forces Key

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Table of Contents Chemistry Practice Test Intermolecular Forces Key

1. Understanding the eBook Chemistry Practice Test Intermolecular Forces Key
 - The Rise of Digital Reading Chemistry Practice Test Intermolecular Forces Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Practice Test Intermolecular Forces Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Practice Test Intermolecular Forces Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Practice Test Intermolecular Forces Key
 - Personalized Recommendations
 - Chemistry Practice Test Intermolecular Forces Key User Reviews and Ratings
 - Chemistry Practice Test Intermolecular Forces Key and Bestseller Lists

5. Accessing Chemistry Practice Test Intermolecular Forces Key Free and Paid eBooks
 - Chemistry Practice Test Intermolecular Forces Key Public Domain eBooks
 - Chemistry Practice Test Intermolecular Forces Key eBook Subscription Services
 - Chemistry Practice Test Intermolecular Forces Key Budget-Friendly Options
6. Navigating Chemistry Practice Test Intermolecular Forces Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemistry Practice Test Intermolecular Forces Key Compatibility with Devices
 - Chemistry Practice Test Intermolecular Forces Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Practice Test Intermolecular Forces Key
 - Highlighting and Note-Taking Chemistry Practice Test Intermolecular Forces Key
 - Interactive Elements Chemistry Practice Test Intermolecular Forces Key
8. Staying Engaged with Chemistry Practice Test Intermolecular Forces Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Practice Test Intermolecular Forces Key
9. Balancing eBooks and Physical Books Chemistry Practice Test Intermolecular Forces Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Practice Test Intermolecular Forces Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chemistry Practice Test Intermolecular Forces Key
 - Setting Reading Goals Chemistry Practice Test Intermolecular Forces Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemistry Practice Test Intermolecular Forces Key
 - Fact-Checking eBook Content of Chemistry Practice Test Intermolecular Forces Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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