

1. Write equilibrium constant expressions for these equations.



2. Use LeChâtelier's principle to predict how each of the following changes would affect this equilibrium.



a. Adding $\text{H}_2\text{O}(g)$ to the system

b. Removing $\text{CO}(g)$ from the system

c. Adding $\text{H}_2(g)$ to the system

d. Adding something to the system to absorb $\text{CO}_2(g)$

3. How would decreasing pressure of the reaction vessel affect these equilibria?



4. How would increasing pressure of the reaction vessel affect these equilibria?



5. How would these equilibria be affected by increasing the temperature?



6. If the equilibrium constant for the following reaction is 2.27×10^{-3} M, what is the equilibrium concentration for hydrogen gas if $[\text{S}_2] = 0.0540$ M and $[\text{H}_2\text{S}] = 0.184$ M?



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JL Elias



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