

## Chapter 19 Study Guide: Acids, Bases &amp; Titration – Answers

1. Differentiate between Arrhenius, Brønsted-Lowry & Lewis acids and bases.

	ACIDS	BASES
Arrhenius	Form hydronium ions ( $\text{H}_3\text{O}^+$ ) in aqueous sol'n	Form hydroxide ions ( $\text{OH}^-$ ) in aqueous sol'n
Brønsted-Lowry	Proton ( $\text{H}^+$ ) donors	Proton ( $\text{H}^+$ ) acceptors
Lewis	Electron pair acceptors	Electron pair donors

2. Label the acid (A), base (B), conjugate acid (CA) and conjugate base (CB) in each of the following reactions:



3. Write the names or formulas for the following acids and bases:

- |  |   |
|--|---|
| a. $\text{H}_2\text{S}$ – hydrosulfuric acid       | e. Nitrous acid – $\text{HNO}_2$                |
| b. $\text{HC}_2\text{H}_3\text{O}_2$ – acetic acid | f. Calcium hydroxide – $\text{Ca}(\text{OH})_2$ |
| c. $\text{NaOH}$ – sodium hydroxide                | g. hydrochloric acid – $\text{HCl}$             |
| d. $\text{HNO}_3$ – nitric acid                    |   |

Use the following formulas for the remaining questions: (these will not be given to you on the test!)

$$K_w = [\text{H}_3\text{O}^+][\text{OH}^-] = 1.0 \times 10^{-14}$$

$$\text{pH} = -\log[\text{H}_3\text{O}^+]$$

$$\text{pOH} = -\log[\text{OH}^-]$$

$$\text{pH} + \text{pOH} = 14$$

$$M_A \cdot V_A \cdot n_A = M_B \cdot V_B \cdot n_B$$

4. What is the pH of
- $1.34 \times 10^{-4}$
- M solution of hydrochloric acid?

$$\text{pH} = -\log [1.34 \times 10^{-4} \text{M}] = 3.873$$

5. Determine the pH and pOH of a
- $1.85 \times 10^{-6}$
- M NaOH solution.

$$\text{pOH} = -\log [1.85 \times 10^{-6} \text{M}] = 5.733$$

$$\text{pH} = 14 - 5.73 = 8.267$$

6. What is the
- $[\text{OH}^-]$
- of an aqueous solution that has a pH of 5.0?

$$\text{pOH} = 9 \quad [\text{OH}^-] = 1 \times 10^{-9} \text{ M}$$

7. What is the pH and the pOH of a solution with a volume of 2000. mL containing 35 grams of NaOH? Is this solution acidic or basic?

$$\frac{35 \text{g NaOH}}{2.000 \text{L}} \times \frac{1 \text{mol NaOH}}{40.00 \text{g NaOH}} = 0.44 \text{ M NaOH}$$

$$\text{pOH} = -\log[0.44 \text{M}] = 0.36 \quad \text{pH} = 14 - 0.36 = 13.64 \quad \text{basic}$$

# Chemistry Acids And Bases Study Guide Answers

**Mary H. Bailey**



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## **Table of Contents Chemistry Acids And Bases Study Guide Answers**

1. Understanding the eBook Chemistry Acids And Bases Study Guide Answers
  - The Rise of Digital Reading Chemistry Acids And Bases Study Guide Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Acids And Bases Study Guide Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in a Chemistry Acids And Bases Study Guide Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Acids And Bases Study Guide Answers
  - Personalized Recommendations
  - Chemistry Acids And Bases Study Guide Answers User Reviews and Ratings

- Chemistry Acids And Bases Study Guide Answers and Bestseller Lists
- 5. Accessing Chemistry Acids And Bases Study Guide Answers Free and Paid eBooks
  - Chemistry Acids And Bases Study Guide Answers Public Domain eBooks
  - Chemistry Acids And Bases Study Guide Answers eBook Subscription Services
  - Chemistry Acids And Bases Study Guide Answers Budget-Friendly Options
- 6. Navigating Chemistry Acids And Bases Study Guide Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Chemistry Acids And Bases Study Guide Answers Compatibility with Devices
  - Chemistry Acids And Bases Study Guide Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chemistry Acids And Bases Study Guide Answers
  - Highlighting and Note-Taking Chemistry Acids And Bases Study Guide Answers
  - Interactive Elements Chemistry Acids And Bases Study Guide Answers
- 8. Staying Engaged with Chemistry Acids And Bases Study Guide Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chemistry Acids And Bases Study Guide Answers
- 9. Balancing eBooks and Physical Books Chemistry Acids And Bases Study Guide Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chemistry Acids And Bases Study Guide Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Chemistry Acids And Bases Study Guide Answers
  - Setting Reading Goals Chemistry Acids And Bases Study Guide Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemistry Acids And Bases Study Guide Answers
  - Fact-Checking eBook Content of Chemistry Acids And Bases Study Guide Answers
  - Distinguishing Credible Sources



13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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