

CHAPTER 8 REVIEW*Chemical Equations and Reactions***SECTION 1****SHORT ANSWER** Answer the following questions in the space provided.

1. Match the symbol on the left with its appropriate description on the right.



(a) A precipitate forms.



(b) A gas forms.



(c) A reversible reaction occurs.



(d) Heat is applied to the reactants.



(e) A chemical is dissolved in water.



(f) A chemical is in the liquid state.

2. Finish balancing the following equation:



3. In each of the following formulas, write the total number of atoms present.

 12 atoms a. 4SO_2 16 atoms b. 8O_2 51 atoms c. $3\text{Al}_2(\text{SO}_4)_3$ 3×10^{24} atoms d. $6 \times 10^{23} \text{HNO}_3$

4. Convert the following word equation into a balanced chemical equation:

aluminum metal + copper(II) fluoride \rightarrow aluminum fluoride + copper metal

5. One way to test the salinity of a water sample is to add a few drops of silver nitrate solution with a known concentration. As the solutions of sodium chloride and silver nitrate mix, a precipitate of silver chloride forms, and sodium nitrate is left in solution. Translate these sentences into a balanced chemical equation.



6. a. Balance the following equation:
- $\text{NaHCO}_3(s) \xrightarrow{\Delta} \text{Na}_2\text{CO}_3(s) + \text{H}_2\text{O}(g) + \text{CO}_2(g)$



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