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(Chapter - 1) (Chemical Reactions and Equations) (Exercises) (Class 10)

Ouestion 14:

In the refining of silver, the recovery of silver from silver nitrate solution involved displacement by copper metal. Write down the reaction involved.

Answer 14:

$$Cu_{(s)}$$
 + $2AgNO_{3(aq)}$ \rightarrow $Cu(NO_3)_{2(aq)}$ + $2Ag_{(s)}$
Copper nitrate Silver

Ouestion 15:

What do you mean by a precipitation reaction? Explain by giving examples.

Answer 15:

A reaction in which an insoluble solid (called precipitate) is formed is called a precipitation reaction.

$$BaCl_2$$
 + K_2SO_4 \longrightarrow $BaSO_4$ + $2KCl$

Barium chloride Potassium sulphate Barium sulphate Potassium chloride

In this reaction, barium sulphate is obtained as a precipitate. Hence, it is a precipitation reaction.

Ouestion 16:

Explain the following in terms of gain or loss of oxygen with two examples each.

(a) Oxidation Answer 16: (b) Reduction

(a) Oxidation: Oxidation is the gain of oxygen.

- $ightharpoonup C + O_2 \longrightarrow CO_2$ $ightharpoonup 2CuO \longrightarrow 2CuO$

Here, carbon (C) is oxidised to carbon dioxide (CO_2) and cuppor is oxidesed to copper (Cu) is oxidesed to copper oxide (CuO).

(b) Reduction: Reduction is the loss of oxygen.

- $\gt CO_2 + H_2 \longrightarrow CO + H_2O$ $\gt CuO + H_2 \longrightarrow Cu + H_2O$

Here, carbon dioxide (CO_2) is reduced to carbon monoxide (CO) and copper oxide (CuO) is reduced to copper (Cu).

Ouestion 17:

A shiny brown coloured element 'X' on heating in air becomes black in colour. Name the element 'X' and the black coloured compound formed.

Answer 17:

The element 'X' is copper (Cu) which is coverted into black coloured copper oxide (CuO).

49.50	1 69	Theat.	20.0
ZCW	+ 02		→ 2CuO
Shiny brown coloured element			Black coloured element

Ouestion 18:

Why do we apply paint on iron articles?

_Answer 18:

Iron articles are painted because applying paints prevents them from rusting. When painted, the contact of iron articles from moisture and air is cut off. Hence, rusting is prevented their presence is essential for rusting to take place.

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