

Class 11 Chemistry Worksheet on Chapter 12 Organic Chemistry – Some Basic Principles & Techniques - Set 3

- Q1. How many pi bonds are there in phosgene?
- (a) One
- (b) Two
- (c) Three
- (d) Zero
- Q2. What is the IUPAC name of CH₃OC₂H₅?
- (a) Ethoxy methane
- (b) Methoxy ethane
- (c) Both (a) and (b)
- (d) None of the above
- Q3. What is the colour of the bromine solution?
- (a) Purple
- (b) Green
- (c) Red
- (d) None of the above
- Q4. Which of the following functional groups is present in RCOOH?
- (a) Alcohol
- (b) Carboxylic acid
- (c) Ketone
- (d) None of the above
- Q5. What is the suffix of esters?
- (a) oate
- (b) ers.
- (c) of
- (d) None of the above
- Q6. Draw the bond-line formula of heptan-4one.
- Q7. What are the hybridisation states of each carbon in CH₃ CH = CH CN?
- Q8. Which of the following carbocation is most stable?
- (CH₃)₃CCH₂*, (CH₃)₃C*, CH₃CH₂CH₂*, CH₃CH*CH₃
- Q9. Draw the resonating structure of phenol and nitrobenzene.

Chemistry Concepts And Applications Chapter 12 Answers

Peter Atkins, Julio de Paula

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Operator Certification Study Guide John Giorgi, 2011-01-12 This book is a revision of the popular study guide for water system last published in 1993 This study resource is a practical tool for treatment plant operators and distribution system personnel as they prepare for the certification exam Actually formatting is used with the sample questions all of which have been reviewed by ABC Association of Board of Certification and are based on information contained in the WSO training series Water Treatment Textbook and the Water Distributor Operation Handbook Math formulas conversation factors and other resource references are also included Previous edition 0 89867 685 1 Merrill Chemistry Robert C. Smoot.Smoot.Richard G. Smith.Iack Price.1998 Instructor's Manual and Solutions to Problems to Accompany Chemistry, a Study of Matter Alfred Benjamin Garrett, William Thomas Lippincott, Frank Henry Verhoek, 1972 **Chemistry** John S. Phillips, Cheryl Wistrom, 2000 **Complete Solutions Guide** Steven S. Zumdahl, Thomas J... Hummel,1998 Solutions Guide for Chemistry Steven S. Zumdahl, Kenneth C. Brooks, 1989 Solutions Guide for Chemistry, Third Edition, by Steven S. Zumdahl Kenneth C. Brooks, 1993 Chemists' Guide to Effective Teaching Norbert J. Pienta, Melanie M. Cooper, Thomas J. Greenbowe, 2005 Part of the Prentice Hall Series in Educational Innovation for Chemistry this unique book is a collection of information examples and references on learning theory teaching methods and pedagogical issues related to teaching chemistry to college students In the last several years there has been considerable activity and research in chemical education and the materials in this book integrate the latest developments in chemistry Each chapter is written by a chemist who has some expertise in the specific technique discussed has done some research on the technique and has applied the technique in a chemistry course **Prentice Hall Physical Science Concepts in** Action Program Planner National Chemistry Physics Earth Science ,2003-11 Prentice Hall Physical Science Concepts in Action helps students make the important connection between the science they read and what they experience every day Relevant content lively explorations and a wealth of hands on activities take students understanding of science beyond the page and into the world around them Now includes even more technology tools and activities to support differentiated Chemistry Principles Student Solutions Guide Steven S. Zumdahl, 2001-08 End of chapter problems with blue instruction numbers appear in this guide These answers also appear at the end of the text These solutions include a mix of odd and even numbered problems comprising about half of the total problem sets found in the text Introduction to General, Organic, and Biological Chemistry Martha J. Gilleland, 1982 Partial Solutions Guide, Chemical Principles, Steven S. Zumdahl, Fifth Edition Thomas J. Hummel, Steven S. Zumdahl, 2004-08 **Elements of Physical Chemistry** Peter Atkins, Julio de Paula, 2013 Elements of Physical Chemistry has been carefully crafted to help students increase their confidence when using physics and mathematics to answer fundamental questions about the structure of molecules how chemical reactions take place and why materials behave the way they do **Essentials of Chemistry** Ralph A. Burns, John William Hill, 1995 The

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