

Задача 11  
Слабые и сред. кислоты

① Дано:

$$C_{\text{HAc}} = 0,01 \text{ м}$$

$$K_a = 7,9 \cdot 10^{-6}$$

$$pH = ?$$

$$pH = -\lg [H^+] = -\lg 2,8 \cdot 10^{-6} = 6 - \lg 2,8 = 6 - 0,45 = 5,55$$

Решение:

Для слабой кислоты:

$$[H^+] = \sqrt{K_a \cdot C_{\text{HAc}}}$$

$$= \sqrt{7,9 \cdot 10^{-6} \cdot 10^{-2}} = 2,8 \cdot 10^{-6} \text{ моль/л}$$

② Дано:

$$K_{\text{PrOH}_2} = 6,3 \cdot 10^{-5}$$

$$pH = ?$$

$$S = \sqrt[3]{\frac{K_{\text{PrOH}_2}}{4}} = \sqrt[3]{\frac{6,3 \cdot 10^{-5}}{4}} = 1,16 \cdot 10^{-2} \text{ моль/л}$$

$$[OH^-] = \frac{S}{2} = \frac{1,16 \cdot 10^{-2}}{2} = 5,8 \cdot 10^{-3} \text{ моль/л}$$

$$[H^+] = \frac{10^{-14}}{[OH^-]} = \frac{10^{-14}}{5,8 \cdot 10^{-3}} = 1,72 \cdot 10^{-12}$$

$$pH = -\lg [H^+] = -\lg 1,72 \cdot 10^{-12} = 12 - \lg 1,72 = 11,76$$

Решение:

$$C_{\text{PrOH}_2} = C_{\text{H}_2\text{Pr}^{2+}} = 2,5 \cdot 10^{-2} \text{ моль/л}$$

$$K_{\text{Pr}} = [Ca^{2+}][OH^-]^2$$

$$= 5 \cdot (2,5)^2 = 31,25$$

③ Дано:

$$V_{\text{HCl}} = V_{\text{NaOH}}$$

$$C_{\text{HCl}} = 0,1 \text{ м}$$

$$V_{\text{NaOH}} = 10 \text{ мл}$$

$$C_{\text{NaOH}} = 0,1 \text{ м}$$

$$K_{\text{PbCl}_2} = 1,6 \cdot 10^{-5}$$

$$PbCl_2 \rightleftharpoons Pb^{2+} + 2Cl^-$$

Решение:

Полное количество:

$$C_{\text{HCl}} = 0,1 \frac{\text{м}}{10} = 0,01 \text{ моль/л}$$

$$C_{\text{NaOH}} = 0,1 \frac{\text{м}}{10} = 0,01 \text{ моль/л}$$

Полное количество ионов:

$$C_{\text{Cl}^-} = C_{\text{HCl}} = 0,01 \text{ моль/л}$$

$$C_{\text{OH}^-} = C_{\text{NaOH}} = 0,01 \text{ моль/л}$$

$$[Pb^{2+}][Cl^-]^2 = 1,6 \cdot 10^{-5} \Rightarrow 10 \cdot 10^{-5} = 10^{-4}$$

# Chemistry Semester 2 Apex

**H Kauffman**



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