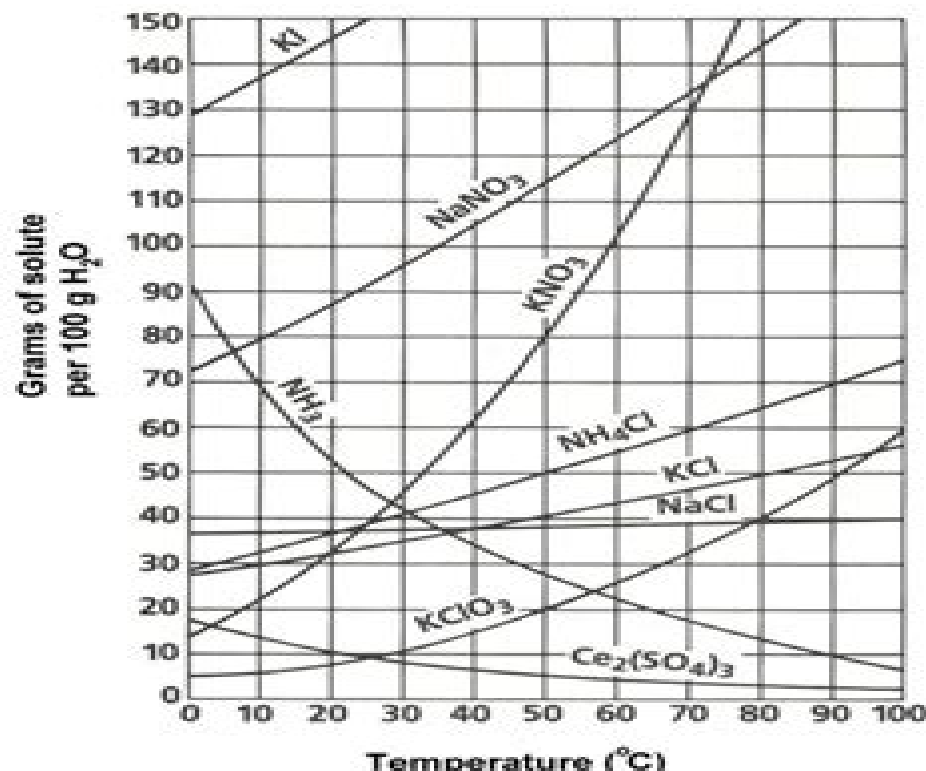


Reading a Solubility Chart

- 1) The **curve** shows the # of grams of solute in a **saturated solution** containing 100 mL or 100 g of water at a certain temperature. (remember: 1 mL H₂O = 1 g H₂O)
- 2) Any amount of solute **below** the line indicates an **unsaturated solution** at a certain temperature
- 3) Any amount of solute **above** the line in which all of the solute has dissolved indicates a **supersaturated solution**.
- 4) If the amount of solute is above the line but has not all dissolved, the solution is saturated and the # grams of solute settled on the bottom of the container = total # g in solution – # g of a saturated solution at that temperature. (according to the curve)
- 5) Solutes whose curves move upward w/ increased temperature are typically solids b/c the solubility of solids increases w/ increased temperature.
- 6) Solutes whose curves move downward w/ increased temperature are typically gases b/c the solubility of gases decreases with increased temperature.



Solubility Problems example (Many may be done by reading the solubility table)

1. At 10°C, 80 g of NaNO₃ will dissolve in 100 mL (a saturated solution)
2. To find the # grams needed to saturate a solution when the volume is NOT 100 mL use the following strategy to find answer: (e. g., 60 mL of H₂O will be saturated with how many grams of NaNO₃ ?)

Start w/ known vol. x Solubility/100mL at set temp. = amount of Solute needed to saturate

Ex. 60 mL H₂O x $\frac{80 \text{ g NaNO}_3}{100 \text{ mL H}_2\text{O}}$ = 48 g NaNO₃ needed to saturate solution

or if the chart is in units of 100 g of H₂O use the density of water conversion 1mL H₂O= 1 g H₂O

Ex. 60 mL H₂O x $\frac{1 \text{ g H}_2\text{O}}{1 \text{ mL H}_2\text{O}}$ x $\frac{80 \text{ g NaNO}_3}{100 \text{ g H}_2\text{O}}$ = 48 g NaNO₃

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 - The Rise of Digital Reading Chemistry Solubility Curve Practice Problems 1 Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Solubility Curve Practice Problems 1 Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Solubility Curve Practice Problems 1 Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Solubility Curve Practice Problems 1 Key
 - Personalized Recommendations
 - Chemistry Solubility Curve Practice Problems 1 Key User Reviews and Ratings

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- 5. Accessing Chemistry Solubility Curve Practice Problems 1 Key Free and Paid eBooks
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 - Highlighting and Note-Taking Chemistry Solubility Curve Practice Problems 1 Key
 - Interactive Elements Chemistry Solubility Curve Practice Problems 1 Key
- 8. Staying Engaged with Chemistry Solubility Curve Practice Problems 1 Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Solubility Curve Practice Problems 1 Key
- 9. Balancing eBooks and Physical Books Chemistry Solubility Curve Practice Problems 1 Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Solubility Curve Practice Problems 1 Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chemistry Solubility Curve Practice Problems 1 Key
 - Setting Reading Goals Chemistry Solubility Curve Practice Problems 1 Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemistry Solubility Curve Practice Problems 1 Key
 - Fact-Checking eBook Content of Chemistry Solubility Curve Practice Problems 1 Key
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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