

## Balancing Nuclear Equations

1. Consider the following nuclear reaction:  $^{232}\text{U}_{92} \rightarrow ^4\text{He}_2 + ^{228}\text{Th}_{90}$

- What type of radiation is produced? \_\_\_\_\_
- How does the number of protons in the reactant compare with the total number of protons in the products? \_\_\_\_\_
- How does the number of neutrons in the reactant compare with the total number of neutrons in the products? \_\_\_\_\_
- How does the reactant's mass number compare with the products' total mass number? \_\_\_\_\_
- Show how each side of the reaction would change if a gamma ray were released.  
\_\_\_\_\_  
\_\_\_\_\_

2. Balance the mass numbers and atomic numbers to complete the equations. Indicate the type of radiation occurring wherever necessary.

a)  $^{131}\text{I}_{53} \rightarrow ^0\text{e}_{-1} + \underline{\hspace{2cm}}$  Type: \_\_\_\_\_

b)  $^{30}\text{P}_{15} \rightarrow ^0\text{e}_{-1} + \underline{\hspace{2cm}}$  Type: \_\_\_\_\_

c)  $^{113}\text{Ag}_{47} \rightarrow \underline{\hspace{2cm}} + \underline{\hspace{2cm}}$  Type: Beta decay

d)  $^{226}\text{Ra}_{88} \rightarrow \underline{\hspace{2cm}} + \underline{\hspace{2cm}} + \underline{\hspace{2cm}}$  Type: Alpha and gamma emission

e)  $^{222}\text{Rn}_{86} \rightarrow ^4\text{He}_2 + \underline{\hspace{2cm}}$  Type: \_\_\_\_\_

f)  $^{129}\text{I}_{53} \rightarrow \underline{\hspace{2cm}} + ^{129}\text{Xe}_{54}$  Type: \_\_\_\_\_

g)  $^{15}\text{O}_8 \rightarrow \underline{\hspace{2cm}} + ^{15}\text{N}_7$  Type: \_\_\_\_\_

h)  $^{239}\text{Pu}_{94} \rightarrow \underline{\hspace{2cm}} + \underline{\hspace{2cm}}$  Type: Alpha decay

# Chemistry Balancing A Nuclear Equation Answers

**Clive Long**



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