1) Given this equation: N2 + 3 H2 ---> 2 NH3, write the following molar ratios:

a) N₂ / H₂

b) N₂/ NH₃

c) H₂ / NH₃

2) Given the following equation: 8 H2 + S8 ---> 8 H2S, write the following molar ratios:

a) H₂ / H₂S

b) H₂ / S₈

c) H₂S / S₈.

3) Answer the following questions for this equation: 2 H2 + O2 ---> 2 H2O

a) What is the H₂ / H₂O molar ratio?

b) Suppose you had 20 moles of H₂ on hand and plenty of O₂, how many moles of H₂O could you make?

c) What is the O₂ / H₂O molar ratio?

d) Suppose you had 20 moles of O2 and enough H2, bow many moles of H2O could you make?

Use this equation: N₂ + 3 H₂ ---> 2 NH₃, for the following problems

a) If you used 1 mole of N₂, how many moles of NH₃ could be produced?

b) If 10 moles of NH₃ were produced, how many moles of N₂ would be required?

c) If 3.00 moles of H2 were used, how many moles of NH3 would be made?

d) If 0.600 moles of NH₃ were produced, how many moles of H₂ are required?

Chemistry Determining Mole Ratios Study Guide Answers

Martha Mackin

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