

Practice Final

Prof. Brindley

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10. What is the molar concentration of chloride ions in a solution prepared by mixing 100.0 mL of 2.0 M KCl with 50.0 mL of a 1.50 M  $\text{CaCl}_2$  solution?

$$\begin{aligned}
 0.1 \text{ L} \cdot \frac{2.0 \text{ mol KCl}}{1 \text{ L}} &= 0.2 \text{ mol Cl}^- \\
 0.05 \text{ L} \cdot \frac{1.5 \text{ mol CaCl}_2}{1 \text{ L}} \cdot \frac{2 \text{ mol Cl}^-}{2 \text{ mol CaCl}_2} &= 0.15 \text{ mol Cl}^- \\
 \hline
 0.2 \text{ mol Cl}^- + 0.15 \text{ mol Cl}^- &= 0.35 \text{ mol Cl}^- \\
 \hline
 \frac{0.35 \text{ mol Cl}^-}{0.15 \text{ L}} &= 2.33 \text{ M Cl}^-
 \end{aligned}$$

11. To what volume should you dilute 133 mL of an 7.90 M  $\text{CuCl}_2$  solution so that 51.5 mL of the diluted solution contains 4.49 g  $\text{CuCl}_2$ ?

$$4.49 \text{ g CuCl}_2 \cdot \frac{1 \text{ mol}}{134.45 \text{ g}} = 0.0334 \text{ mol CuCl}_2$$

$$M_1 V_1 = M_2 V_2 \rightarrow 7.90 \text{ M} \cdot 0.133 \text{ L} = 0.648 \text{ M} \cdot V_2$$

$$V_2 = \frac{7.9 \cdot 0.133}{0.648} = 1620 \text{ mL}$$

Dilute to 1620 mL

Add 1620 - 133 mL of water = 1487 mL

12. In lab you performed a titration of acetic acid with sodium hydroxide. 83 mL of sodium hydroxide at a concentration of 0.015 M were needed to neutralize 58.2 mL of an acetic acid solution. What is the concentration of the acetic acid solution?

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**Jicheng Xie**



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