

EQUILIBRIUM CONSTANT - WORKSHEET 2_ANSWERS

1. Write the equilibrium constant expression for each of the following reactions:



$$K = \frac{[\text{NO}_2]^2}{[\text{N}_2][\text{O}_2]}$$



$$K = \frac{[\text{C}_2\text{H}_6]}{[\text{H}_2]^3}$$



$$K = \frac{1}{[\text{I}_2(\text{aq})]}$$

2. For the following reaction:



At equilibrium, $[\text{Fe}^{2+}] = 0.50\text{ M}$, $[\text{Ag}^+] = 1.0\text{ M}$, $[\text{Fe}^{3+}] = 1.50\text{ M}$

- a) Calculate the equilibrium constant for the reaction.

$$K = 3.0$$

- b) If at equilibrium, $[\text{Fe}^{2+}] = 0.20\text{ M}$ and $[\text{Ag}^+] = 0.30\text{ M}$, what must be the concentration of Fe^{3+} ?

$$[\text{Fe}^{3+}] = 0.18\text{ mol/L}$$

3. $2\text{NO}(\text{g}) + \text{O}_2(\text{g}) \rightleftharpoons 2\text{NO}_2(\text{g})$ $K = 5.0 \times 10^{-5}$

Calculate the equilibrium concentration of oxygen gas if the NO and NO_2 concentrations are equal at equilibrium.

$$[\text{O}_2] = 2.0 \times 10^{-6}\text{ mol/L}$$

4. a) $\text{CH}_3\text{COOH}(\text{aq}) \rightleftharpoons \text{H}^+(\text{aq}) + \text{CH}_3\text{COO}^-(\text{aq})$ $K = 1.8 \times 10^{-5}$



Which of the above reactions would have the greatest amount of products present at equilibrium compared to reactants?

Reaction b. Since K is larger the amount of products must be larger, because K is products/reactants.

5. For the reaction:



at an equilibrium temperature of 1000°C , a 1.00 L flask contains 0.120 moles of ammonia, 1.03 moles of nitrogen and 1.62 moles of hydrogen. Calculate the equilibrium constant at this temperature.

$$K = 3.29 \times 10^{-2}$$

Chemistry Equilibrium Practice Problems With Answers

**Patrick E. McMahon, Rosemary
McMahon, Bohdan Khomtchouk**



Chemistry Equilibrium Practice Problems With Answers:

Survival Guide to General Chemistry Patrick E. McMahon, Rosemary McMahon, Bohdan Khomtchouk, 2019-02-13 This work evolved over thirty combined years of teaching general chemistry to a variety of student demographics The focus is not to recap or review the theoretical concepts well described in the available texts Instead the topics and descriptions in this book make available specific detailed step by step methods and procedures for solving the major types of problems in general chemistry Explanations instructional process sequences solved examples and completely solved practice problems are greatly expanded containing significantly more detail than can usually be devoted to in a comprehensive text Many chapters also provide alternative viewpoints as an aid to understanding Key Features The authors have included every major topic in the first semester of general chemistry and most major topics from the second semester Each is written in a specific and detailed step by step process for problem solving whether mathematical or conceptual Each topic has greatly expanded examples and solved practice problems containing significantly more detail than found in comprehensive texts Includes a chapter designed to eliminate confusion concerning acid base reactions which often persists through working with acid base equilibrium Many chapters provide alternative viewpoints as an aid to understanding This book addresses a very real need for a large number of incoming freshman in STEM fields *Physical Chemistry Student Solutions Manual* Charles Trapp, Peter Atkins, Marshall Cady, 2006-08-11 Change 21 *Questions & Answers About Block Scheduling* John Brucato, 2014-04-11 For administrators and others involved in the transition to block schedules this book provides answers to the complex and challenging questions raised by the curious and the skeptical It demonstrates how to overcome obstacles to systemic school improvements

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