

Reactions of Amines and Epoxides

Preparation of Amines

Amination of alcohols: Polyalkylation is a serious problem for indirect synthesis of amines



A summary of Ring-Opening Reactions of Epoxides with Strong and Weak Nucleophiles



Bond Types, Strength, Length, Polarity and Dipole Moment

Two types of covalent bond

There are two types of covalent bond: ionic bond and covalent bond

- Ionic bonds are formed by the transfer of electrons from one element to another (ionization potential) and electron affinity (electron affinity).



- Covalent bonds are formed by sharing of electrons between two atoms.



The Relative Electronegativity of Atoms



Bond and Bonding in a Molecule



Bonding in a Molecule



There are two types of bond

Definition of ionic bond



Definition of covalent bond



Definition of ionic bond



Definition of covalent bond

Definition of ionic bond

Sigma and pi bonds

Covalent bonds are classified into sigma and pi bonds

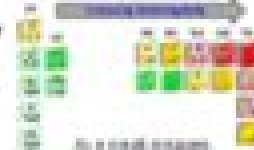
- Sigma bonds are formed by head-on overlap of orbitals.
- pi bonds are formed by side-on overlap of orbitals.
- A single bond is a sigma bond.
- A double bond consists of one sigma and one pi bond.



pi bonds are stronger than sigma bonds

Polar and Nonpolar Covalent Bonds

Depending on the difference in the electronegativity values, covalent bonds are divided into polar and nonpolar.



- If the difference in electronegativity is less than 0.5, the electrons are shared equally (nonpolar bond). For example, C-C and C-H bonds.
- If the difference in electronegativity is between 0.5 and 1.7, the electrons are shared unequally (polar bond). For example, C-O and C-N bonds.
- A difference of 1.7 or higher is ionic (the electrons are not shared).

Dipole Moment

The polarity of a bond is described by the dipole moment (μ).



- The dipole moment is a vector pointing from the less electronegative atom to the more electronegative atom.
- The vector dipole moment is a product of the charge (q) and the distance (r) between the charges ($\mu = qr$).
- μ (Debye) indicates the electron distribution.
- μ (Debye) indicates the electron distribution.

Molecular Dipole

To determine the molecular polarity, the vector sum of the dipole moments

The polarity of the molecule is determined by the vector sum of the dipole moments. A good example is carbon dioxide.



Carbon dioxide is a nonpolar molecule.

Bond Length and Bond Strength

The bond length is the distance between the nuclei of the two atoms.

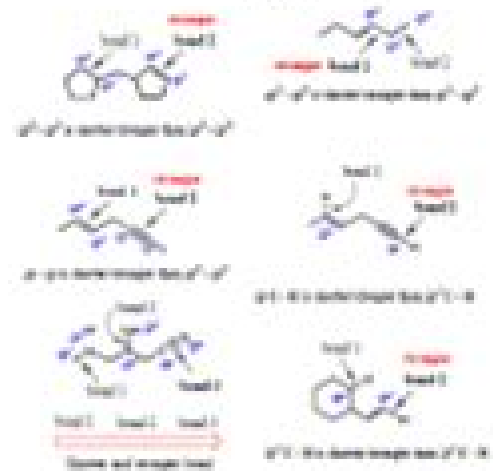


The bond length is the distance between the nuclei of the two atoms.

Average values of C-C and C-H bonds depending on hybridization			
C-C Bond	sp^3-sp^3	sp^2-sp^2	$sp-sp$
Hybridization	sp^3-sp^3	sp^2-sp^2	$sp-sp$
Length	1.54 Å	1.34 Å	1.20 Å
Strength	33 kcal/mol	112 kcal/mol	200 kcal/mol
Hybridization	sp^3-sp^3	sp^2-sp^2	$sp-sp$
C-H Bond	sp^3-sp^3	sp^2-sp^2	$sp-sp$
Length	1.09 Å	1.08 Å	1.06 Å
Strength	98 kcal/mol	110 kcal/mol	120 kcal/mol

Higher s character means shorter and stronger bond

Example



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