

- (c) Name the gas evolved in each of the following cases: [5]
- (i) Copper carbonate is heated strongly.
 - (ii) Action of dilute hydrochloric acid on sodium sulphite.
 - (iii) Nitrogen combines with hydrogen.
 - (iv) Action of dilute sulphuric acid on sodium carbonate.
 - (v) Addition of sodium to cold water.
- (d) Give a reason for each of the following: [5]
- (i) Noble gases do not combine with other elements to form molecules.
 - (ii) Dilute nitric acid cannot be used in the preparation of hydrogen.
 - (iii) Hydrogen although lighter than air, is not collected by the downward displacement of air.
 - (iv) Anhydrous calcium chloride turns into a colourless solution when exposed to air.
 - (v) Metals form positive ions.
- (e) What do you observe when: [5]
- (i) Iodine crystals are heated in a test tube?
 - (ii) Iron nails are added to copper sulphate solution?
 - (iii) Lead nitrate crystals are heated strongly?
 - (iv) Ferric chloride crystals are exposed to atmosphere for some time?
 - (v) Blue copper sulphate crystals are heated?
- (f) (i) State Boyle's law. [5]
- (ii) What is standard temperature and pressure?
- (iii) Calculate the percentage of nitrogen in ammonium nitrate. $[\text{NH}_4\text{NO}_3]$
[N = 14, H = 1, O = 16]

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