

## Electron Configuration Practice Worksheet

In the space below, write the unabbreviated electron configurations of the following elements:

- 1) sodium  $1s^2 2s^2 2p^6 3s^1$
- 2) iron  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^6$
- 3) bromine  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^5$
- 4) barium  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2$
- 5) neptunium  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2 4f^{14} 5d^{10} 6p^6 7s^2 5f^4$   
93e

In the space below, write the abbreviated electron configurations of the following elements:

- 6) cobalt  $[Ar] 4s^2 3d^7$
- 7) silver  $[Kr] 5s^2 4d^9$   $\rightarrow [Kr] 5s^1 4d^{10}$  (exception  $\rightarrow$  more stable)
- 8) tellurium  $[Kr] 5s^2 4d^{10} 5p^4$
- 9) radium  $[Rn] 7s^2$
- 10) lawrencium  $[Rn] 7s^2 5f^{14} 6d^1$   $\checkmark$  last electron in d of shell

Determine what elements are denoted by the following electron configurations:

- 11)  $1s^2 2s^2 2p^6 3s^2 3p^4$  Sulfur
- 12)  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^1$  Rubidium
- 13)  $[Kr] 5s^2 4d^{10} 5p^3$  Antimony
- 14)  $[Xe] 6s^2 4f^{14} 5d^6$  Osmium
- 15)  $[Rn] 7s^2 5f^{11}$  Terbium

Determine which of the following electron configurations are not valid:

- 16)  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 4d^{10} 4p^5$  not valid (3d not there)
- 17)  $1s^2 2s^2 2p^6 3s^2 3d^5$  not valid (3d comes after 3p)
- 18)  $[Ra] 7s^2 5f^6$  not valid [Rn] is not a noble gas
- 19)  $[Kr] 5s^2 4d^{10} 5p^5$  valid (iodine)
- 20)  $[Xe]$  not valid (noble gases cannot be abbreviated)

# Configuration Practice Skill Practice 13

**Kerry E. Kilpatrick**



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## **Table of Contents Configuration Practice Skill Practice 13**

1. Understanding the eBook Configuration Practice Skill Practice 13
  - The Rise of Digital Reading Configuration Practice Skill Practice 13
  - Advantages of eBooks Over Traditional Books
2. Identifying Configuration Practice Skill Practice 13
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Configuration Practice Skill Practice 13
  - User-Friendly Interface
4. Exploring eBook Recommendations from Configuration Practice Skill Practice 13
  - Personalized Recommendations
  - Configuration Practice Skill Practice 13 User Reviews and Ratings
  - Configuration Practice Skill Practice 13 and Bestseller Lists

5. Accessing Configuration Practice Skill Practice 13 Free and Paid eBooks
  - Configuration Practice Skill Practice 13 Public Domain eBooks
  - Configuration Practice Skill Practice 13 eBook Subscription Services
  - Configuration Practice Skill Practice 13 Budget-Friendly Options
6. Navigating Configuration Practice Skill Practice 13 eBook Formats
  - ePub, PDF, MOBI, and More
  - Configuration Practice Skill Practice 13 Compatibility with Devices
  - Configuration Practice Skill Practice 13 Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Configuration Practice Skill Practice 13
  - Highlighting and Note-Taking Configuration Practice Skill Practice 13
  - Interactive Elements Configuration Practice Skill Practice 13
8. Staying Engaged with Configuration Practice Skill Practice 13
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Configuration Practice Skill Practice 13
9. Balancing eBooks and Physical Books Configuration Practice Skill Practice 13
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Configuration Practice Skill Practice 13
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Configuration Practice Skill Practice 13
  - Setting Reading Goals Configuration Practice Skill Practice 13
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Configuration Practice Skill Practice 13
  - Fact-Checking eBook Content of Configuration Practice Skill Practice 13
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

#### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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