

Don't Flip Your Lid: Comparing Intermolecular Forces (from Laying the Foundation)

Data and Observations

The order of melting (from 1st to last): paraffin, iodine, glucose. Sodium chloride does not melt.

Analysis

From what you know about intermolecular forces, explain the relative melting points.

Paraffin is a medium sized nonpolar molecule and iodine is a larger nonpolar molecule. Both paraffin and iodine only contain London forces. However, paraffin has a lower melting point, which makes it melt first. The lower melting point is due to the smaller size of the paraffin molecule. Since iodine is larger, it has more electrons and a larger, more polarizable electron cloud.

Glucose is a medium sized polar molecule that contains London forces. Polar molecules can also form hydrogen bonds. This creates stronger intermolecular forces, producing higher melting points.

Sodium chloride did not melt because it contains ionic bonds, which are very strong. This strength is due to the electrostatic forces within the crystal lattice, which goes in all directions.

Conclusion

1. Sucrose would have a higher melting point than glucose because it has more electrons. This makes its electron cloud, which is larger, more polarizable. The more polarizable, the more the electron cloud can be distorted, producing a temporary dipole. The larger the dipole created from the polarizability, the larger the intermolecular forces. Larger intermolecular forces mean a higher melting point.
2. Both water and dihydrogen sulfide are bent, polar molecules. However, water has a hydrogen bond between the hydrogen and a very electronegative atom, oxygen. Dihydrogen sulfide do not have these hydrogen bonds, since they only form between hydrogen and electronegative atoms such as oxygen, nitrogen or fluorine.



3. All halogens contain London forces (or dispersion forces) only. London dispersion forces depend on the number of electrons and the size of the electron cloud (greater polarizability). The more electrons a molecule has, the greater the intermolecular forces. The greater the molecular forces, the stronger the attraction between molecules. Iodine has the largest number of electrons, therefore a stronger attraction, creating a solid. Fluorine and chlorine have fewer electrons, creating weaker forces and forming a gas. Bromine is somewhere in between. It's forces are stronger than those in fluorine and chlorine, but not as strong as in iodine.

Note: Just because a molecule is bigger (more massive) does not mean it has greater intermolecular forces. This argument confuses gravitational attraction with London forces. Gravitational attraction depends on mass and is very weak and play little to no part in the attraction between molecules. It is the electrostatic forces between molecules due to the distortion of the electron cloud that determines the attraction between molecules.

Dont Flip Your Lid Comparing Intermolecular Forces

Richard Bailey



Dont Flip Your Lid Comparing Intermolecular Forces:

Decoding **Dont Flip Your Lid Comparing Intermolecular Forces**: Revealing the Captivating Potential of Verbal Expression

In a time characterized by interconnectedness and an insatiable thirst for knowledge, the captivating potential of verbal expression has emerged as a formidable force. Its power to evoke sentiments, stimulate introspection, and incite profound transformations is genuinely awe-inspiring. Within the pages of "**Dont Flip Your Lid Comparing Intermolecular Forces**," a mesmerizing literary creation penned by a celebrated wordsmith, readers set about an enlightening odyssey, unraveling the intricate significance of language and its enduring effect on our lives. In this appraisal, we shall explore the book's central themes, evaluate its distinctive writing style, and gauge its pervasive influence on the hearts and minds of its readership.

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