

DRAWING COVALENT BONDS

Introduction:

A **covalent bond** forms between **nonmetals**. Covalent bonds involve **sharing pairs of electrons** between two atoms so that each atom gains a full outer valence shell. Sharing 1 pair of electrons creates a **single covalent bond**, sharing 2 pairs of electrons creates a **double covalent bond**, and sharing 3 pairs of electrons creates a **triple covalent bond**.



How to Draw Simple Covalent Bond Diagrams	
1.	Draw the electron dot diagrams for each element. <i>HINT: This will help you see the number of unpaired valence electrons for each element.</i>
2.	Count the total number of valence electrons between all atoms. <i>HINT: The total number of valence electrons should be 8 in most cases, except 2 hydrogens. If there are not 8, add more atoms of the element with the least number of valence electrons to start.</i>
3.	Determine which element is the central atom in the molecule, if needed. <i>HINT: The central atom is typically the atom with the largest valence number or the atom that there is only one of in the formula. If there are only 2 atoms in the formula, there is no central atom!</i>
4.	Pair up single electrons from different atoms to form covalent bonds and circle each pair. <i>HINT: Keep doing this until each single electron from all atoms have been paired up. You may need to use double or triple bonds to make this work. This step takes trial and error. If one arrangement is not working, try adding another element or trying a new structure. Practice makes perfect!</i>

Activity: Drawing Covalent Bond Diagrams:

- For each covalent compound, you have been given the element names and the chemical formula.
- Follow the steps listed above to create a covalent bonding diagram for the compound.
- Remember, this takes trial and error. The first 2 have been done for you as examples.

#	Nonmetal	Nonmetal	Formula	Bonding Diagram
A.	hydrogen $\text{H} \cdot$ <i>(must share 1)</i>	chlorine $\cdot \ddot{\text{Cl}} \cdot$ <i>(must share 1)</i>	HCl <i>(Formula shows 1 H and 1 Cl)</i>	
B.	hydrogen $\text{H} \cdot$ <i>(must share 1)</i>	oxygen $\cdot \ddot{\text{O}} \cdot$ <i>(must share 2)</i>	H_2O <i>(Formula shows 2 H and 1 O)</i>	

Drawing Covalent Bonds Worksheet

Henry Sinclair Pickering



Drawing Covalent Bonds Worksheet:

Concept-Based Curriculum and Instruction for the Thinking Classroom H. Lynn Erickson, Lois A. Lanning, Rachel French, 2017-02-02 Develop students critical thinking abstract reasoning and creative learning skills with concept based teaching Take learning beyond the facts with a teaching approach that develops conceptual thinking and problem solving skills A Concept Based curriculum recaptures students innate curiosity about the world and provides the thrilling feeling of using one's mind well Concept Based teachers will learn how to Meet the demands of rigorous academic standards Use the Structure of Knowledge and Process when designing disciplinary units Engage students in inquiry through inductive teaching Identify conceptual lenses and craft quality generalizations Using Multimodal Representations to Support Learning in the Science Classroom Brian Hand, Mark McDermott, Vaughan Prain, 2015-11-06 This book provides an international perspective of current work aimed at both clarifying the theoretical foundations for the use of multimodal representations as a part of effective science education pedagogy and the pragmatic application of research findings to actual classroom settings Intended for a wide ranging audience from science education faculty members and researchers to classroom teachers school administrators and curriculum developers the studies reported in this book can inform best practices in K 12 classrooms of all science disciplines and provide models of how to improve science literacy for all students Specific descriptions of classroom activities aimed at helping infuses the use of multimodal representations in classrooms are combined with discussion of the impact on student learning Overarching findings from a synthesis of the various studies are presented to help assert appropriate pedagogical and instructional implications as well as to suggest further avenues of research

CLIMATOLOGY NARAYAN CHANDER, 2024-02-20 Note Anyone can request the PDF version of this practice set workbook by emailing me at cbsenet4u gmail com I will send you a PDF version of this workbook This book has been designed for candidates preparing for various competitive examinations It contains many objective questions specifically designed for different exams Answer keys are provided at the end of each page It will undoubtedly serve as the best preparation material for aspirants This book is an engaging quiz eBook for all and offers something for everyone This book will satisfy the curiosity of most students while also challenging their trivia skills and introducing them to new information Use this invaluable book to test your subject matter expertise Multiple choice exams are a common assessment method that all prospective candidates must be familiar with in today's academic environment Although the majority of students are accustomed to this MCQ format many are not well versed in it To achieve success in MCQ tests quizzes and trivia challenges one requires test taking techniques and skills in addition to subject knowledge It also provides you with the skills and information you need to achieve a good score in challenging tests or competitive examinations Whether you have studied the subject on your own read for pleasure or completed coursework it will assess your knowledge and prepare you for competitive exams quizzes trivia and more *Holt Chemistry* Ralph Thomas Myers, 2004 Chemical Misconceptions Keith

Taber, 2002 Part one includes information on some of the key alternative conceptions that have been uncovered by research and general ideas for helping students with the development of scientific conceptions **Prentice Hall Physical Science Concepts in Action Program Planner National Chemistry Physics Earth Science** ,2003-11 Prentice Hall Physical Science Concepts in Action helps students make the important connection between the science they read and what they experience every day Relevant content lively explorations and a wealth of hands on activities take students understanding of science beyond the page and into the world around them Now includes even more technology tools and activities to support differentiated instruction Discover Science: Teacher's annotated edition ,1991 Science content helps develop the skills needed to understand how science works learn new concepts solve problems and make decisions in today's technological society **Atoms, Electrons, Structure and Bonding** Primrose Kitten, 2017-06-19 I'm constantly telling you the best way to learn is by practicing questions so I've made you a book full of practice questions Multiple choice questions to reflect the style of exam questions activities to complete equations for you to balance compounds for you to work out the formula for lots of things that you need to recall and practice long answer exam style questions This book is not designed as a text book or revision guide but as a workbook There are lots of good and bad expensive and free revision guides out there on my YouTube channel and other great websites So there is no point in me adding to the masses All the teaching all the new content is available for free on my YouTube channel this book is for you to practice and learn The best way to approach this is to watch the teaching video and make notes or after class try a section and check the answers Any corrections that are needed after the book is published will be listed on my website www.primrosekitten.com these will be corrected in the next version of the book Answers are provided for the sections where you need to work out the answers for yourself not the sections where you are just filling things in from a video or website **Atoms Electrons Structure and Bonding Workbook** Topics Covered are Some of this content has also been published in the Summer Start to A Level Chemistry and a practice exam paper **Atomic Structure** 20 Multiple choice questions **Properties of Ionic Compounds** 15 multiple choice questions **Reference table of common ions formulae** **Formula of Ionic Compounds** 65 formulas to work out **Drawing Ionic Bonding** 10 Compounds **Simple Covalent Bonding** 20 multiple choice questions **Drawing Covalent Bonding** 10 Compounds **Summary Table for the 4 Different Types of Bonding** **Electron configurations** **Drawing electron configurations** **Drawing Electron Configurations** **Spot the mistake Electronic Configuration** 20 multiple choice questions **Exceptions to the Octet Rule** **Oxidation Numbers** 20 multiple choice questions **Balancing Equations using the oxidation numbers method** 20 to practice **Salt Equations** 20 equations to complete and balance **Shapes of Molecules Investigation** **Shapes of Molecules and Bond Angles** 20 multiple choice questions **Electronegativity and Bond Polarity Investigation** **Intermolecular Bonding** 10 multiple choice questions **Electrons Structure and Bonding Practice Exam Questions** **Answers** *The Covalent Bond* Henry Sinclair Pickering, 1977 *Chemistry of the Covalent Bond* Leallyn B. Clapp, 1950 **Understanding Chemistry: Chemical bonding** ,1967 Chemistry of the

Covalent Bond Leallyn Burr Clapp,1957 *Polar Covalence* R Sanderson,2012-12-02 Polar Covalence provides a detailed account of a successful approach to understanding chemistry from knowledge of atomic structure and the properties that result from this structure This book discusses the nature of multiple bonds Organized into 16 chapters this book begins with an overview of the interrelationships of various basic atomic properties This text then describes chemical bonding which can only occur when the nuclei of both atoms can attract the same electrons Other chapters consider the bond energy of multiple bonds which can be determined by calculating the energy in the usual way as though the bonds were single but of the experimental length This book discusses as well the reduction of the lone pair bond weakening effect through the formation of multiple bonds The final chapter deals with the relative roles of principles and practice in the teaching of inorganic and general chemistry This book is a valuable resource for chemists and students Chemical Bonding M.S. Sethi & M. Satake,2010 Contents Chemical Bonding I Basic Concepts Chemical Bonding II Additional Aspects Intermolecular Force and Crystal Structures *Let's Bond Together! Explaining Why Atoms Bond, Types of Bonding and Electron Dot Diagrams | Grade 6-8 Physical Science* Dot EDU,2024-04-15 Dive into the captivating world of chemical bonds with this essential guide for grades 6 8 a must have for the US STEM curriculum This book demystifies the intricate dance of atoms as they bond to form the myriad substances that make up our universe From the basics of covalent and ionic bonds to the detailed exploration of electron dot diagrams it s an invaluable resource for teachers homeschooling parents and librarians Discover the building blocks of matter and how they come together in complex ways Add this to your library to spark a love of science in young learners Teaching Chemical Bonding Margaret Irene Lindsay,1995 This document presents an instructional strategy for teaching chemical bonding using parables and music Games student interactions and worksheets are included in the lesson plans Topics include metallic bonding covalent bonding including molecular and network structure and ionic bonding JRH **Structure and Bonding** Jack Barrett,2002-05-09 Structure and Bonding covers introductory atomic and molecular theory as given in first and second year undergraduate courses at university level This book explains in non mathematical terms where possible the factors that govern covalent bond formation the lengths and strengths of bonds and molecular shapes Throughout the book theoretical concepts and experimental evidence are integrated An introductory chapter summarizes the principles on which the Periodic Table is established and describes the periodicity of various atomic properties which are relevant to chemical bonding Symmetry and group theory are introduced to serve as the basis of all molecular orbital treatments of molecules This basis is then applied to a variety of covalent molecules with discussions of bond lengths and angles and hence molecular shapes Extensive comparisons of valence bond theory and VSEPR theory with molecular orbital theory are included Metallic bonding is related to electrical conduction and semi conduction The energetics of ionic bond formation and the transition from ionic to covalent bonding is also covered

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